

Rhodium Complexes in P–C Bond Formation: Key Role of a Hydrido Ligand

Víctor Varela-Izquierdo,[†] Ana M. Geer,[†] Janeth Navarro,[†] José A. López, Miguel A. Ciriano, and Cristina Tejel^{*}



Cite This: <https://dx.doi.org/10.1021/jacs.0c11010>



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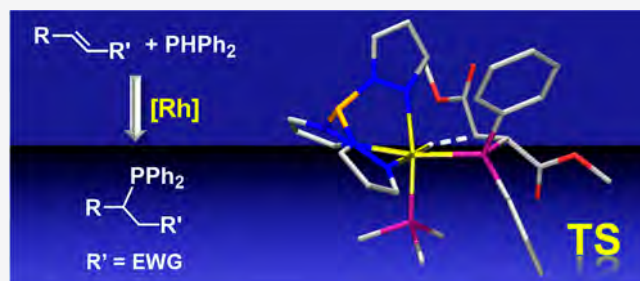


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ABSTRACT: Olefin hydrophosphanation is an attractive route for the atom-economical synthesis of functionalized phosphanes. This reaction involves the formation of P–C and H–C bonds. Thus, complexes that contain both hydrido and phosphanido functionalities are of great interest for the development of effective and fast catalysts. Herein, we showcase the excellent activity of one of them, $[\text{Rh}(\text{Tp})\text{H}(\text{PMe}_3)(\text{PPh}_2)]$ (**1**), in the hydrophosphanation of a wide range of olefins. In addition to the required nucleophilicity of the phosphanido moiety to accomplish the P–C bond formation, the key role of the hydride ligand in **1** has been disclosed by both experimental results and DFT calculations. An additional Rh–H...C stabilization in some intermediates or transition states favors the hydrogen transfer reaction from rhodium to carbon to form the H–C bond. Further support for our proposal arises from the poor activity exhibited by the related chloride complex $[\text{Rh}(\text{Tp})\text{Cl}(\text{PMe}_3)(\text{PPh}_2)]$ as well as from stoichiometric and kinetic studies.



INTRODUCTION

Carbon–phosphorus (C–P) bond formation is a key step in the synthesis of organophosphanes, which play essential roles as ligands for transition metal complexes and in organocatalysis as building blocks for medicinal and supramolecular chemistry and have extensive commercial applications.¹ However, classical methodologies for the synthesis of organophosphanes suffer from safety concerns (handling corrosive and flammable compounds), use of protective groups that bring in additional synthetic steps, lack of selectivity, and limited scope.^{1d,2} Therefore, there has been a growing interest in the development of more efficient synthetic strategies, among which the net addition of a P–H bond to unsaturated substrates is one of the most attractive due to the safety, selectivity, and 100% atom-economy provided by this approach.³ In particular, hydrophosphanation of alkenes/alkynes is a powerful synthetic tool to create a plethora of compounds containing a C–P bond such as secondary or tertiary phosphanes with new alkyl substituents.⁴

Among the wide range of catalysts with a M–PR₂ moiety reported for this reaction,⁵ our attention has focused on those arising from the oxidative-addition of the P–H bond. This activation reaction eventually gives hydrido-phosphanido intermediates, allowing thus the entry of the phosphanido functionality into the coordination sphere of the metal in a simple and clean way.

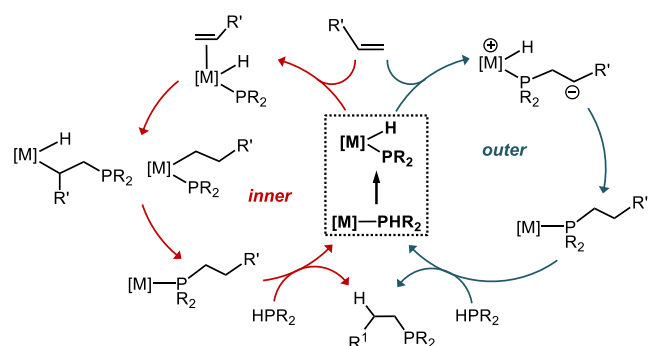
From a mechanistic point of view, these intermediates can follow two main alternatives for the next P–C bond formation

step: either coordination of the alkene/alkyne followed by insertion into the M–P⁶ or the M–H⁷ bonds (inner-sphere pathway) or outer-sphere phospho-Michael additions.⁸ A further reductive-elimination reaction in the former case and a hydrido transfer to the intermediate carbanion in the latter one render phosphane complexes, in which a simple phosphane exchange closes the catalytic cycles (Scheme 1).

Rhodium complexes have been scarcely studied in this field. Particularly pertinent to the present work are the rhodium(III) hydrido-phosphanido intermediates observed by Tilley, which are suitable for phosphorus–sulfur coupling,⁹ and a rhodium(V) bis(hydrido-phosphanido) complex proposed by Brookhart in the catalytic dehydrocoupling of secondary phosphanes.¹⁰ More recently, N-heterocyclic carbene (NHC) rhodium(I) complexes have been shown to be efficient catalysts for the double hydrophosphanation of alkynes,¹¹ whereas we have shown the influence of the ancillary ligands on the formation of mononuclear hydrido-phosphanido rhodium complexes¹² and confirmed that some of them are intermediates in dehydrocoupling reactions.¹³

Received: October 19, 2020

Scheme 1. Proposed Catalytic Cycles for Olefin Hydrophosphanation Involving Hydrido-phosphanido Complexes^a



^aR' = electron withdrawing group (EWG) for the cycle in blue.

Herein, we showcase a full study on the catalytic activity of a hydrido-phosphanido rhodium complex in the hydrophosphanation of alkenes with diphenylphosphane. We disclose the key

role of the hydride ligand in providing a low energy pathway for the C–H bond formation after the critical P–C bond formation, supported by kinetics and in agreement with DFT studies. This key step seems to be the origin of its exceptionally good activity in olefin hydrophosphanation.

RESULTS AND DISCUSSION

Catalytic Assays. The hydrido-phosphanido complex [Rh(Tp)H(PMe₃)(PPh₂)] (**1**)¹² was examined as a catalyst for the hydrophosphanation with PPh₂ of the olefins listed in Table 1. All the reactions were fully selective to the *anti*-Markovnikov product, as typically observed for hydrophosphanation reactions, with no traces of the Markovnikov isomer detected by NMR spectroscopy.

Complex **1** was found to be extremely active in the hydrophosphanation of methylacrylate (MeAC), with a full conversion in less than 5 min at room temperature (rt) even with very low catalyst loadings (0.1% mol cat., entry 2). Compared with other catalysts reported in the literature,^{4e,g,l,v} complex **1** is, to the best of our knowledge, the fastest catalyst for the hydrophosphanation of methylacrylate with PPh₂.

Table 1. Synthesis of Functionalized Phosphanes Catalyzed by Complex **1**^a

Entry	Alkene	Product ^c	1 (mol%)	Time (min)	Conv. (%) ^b	Isolated yield (%)
1	Methylacrylate		5	< 5	> 99	98
2	Methylacrylate		0.1	< 5	> 99	
3	Methylacrylate		0.02	669	98	
4	Acrylonitrile		5	18	> 99	97
5	Acrylonitrile		1	116	95	
6	Dimethylfumarate		5	35	> 99	98
7	Dimethylfumarate		1.7	46	95	
8	Dimethylmaleate		5	50	> 99	98
9	Methylmethacrylate		5	568	86	98
10	Cinnamaldehyde		5	25	86	72
11	Crotonaldehyde		5	43	66	76
12	Benzylideneacetone		5	320	95	89
13	Cyclohexenone		3.8	269	96	82

^aReaction conditions: PPh₂ and alkene (0.48 mmol), **1** (0.024 mmol, 5% mol) in 0.7 mL of C₆D₆ at room temperature. For catalytic loadings of 1% or lower, the appropriate amount of catalyst was taken from a stock solution and the total volume was adjusted to 0.7 mL. [PPh₂] = [olefin] = 2.71 (entry 2), 0.74 (entry 3), 1.23 (entry 5), 0.99 (entry 7), 0.99 (entry 13), M. ^bDetermined by ¹H NMR spectroscopy. ^cThe original position of the C=C bond in the alkene is marked in red in the product.

Nonetheless, $\text{Ph}_2\text{PhCH}_2\text{CH}_2\text{CO}_2\text{Me}$ can be alternatively prepared in 1 h by the simple use of hydrated tetraethylammonium hydroxide.¹⁴

The related acrylonitrile (ACN), with a more electron withdrawing group (EWG) than methylacrylate (MeAC),¹⁵ was also hydrophosphanated with PPh_2 . The catalysis was found to be slower than for MeAC, requiring 18 min at rt; higher catalyst loadings (5% mol cat, entry 4) and around 2 h if more diluted catalytic solutions were used (1% mol cat, entry 5). These results were quite unexpected, since in a classic outer-sphere phospho-Michael addition, a more EWG at the α -position of the olefin is expected to produce an increase in the reaction rate. Despite this, complex **1** operates more rapidly and with higher conversions in the hydrophosphanation of acrylonitrile with PPh_2 than other catalysts reported in the literature. As a matter of fact, a TOF of 30 h^{-1} has been reported by Morris¹⁶ with $[\text{Ru}(\text{Cp}^*)(\text{PPh}_2)(\text{Ph}_2\text{PCH}=\text{CHPPH}_2)]$ as catalyst (1% mol, rt), whereas values of 10 and 2 h^{-1} were observed by Glueck^{6b} with $[\text{Pt}(\text{dppe})(\text{H}_2\text{C}=\text{CHCN})]$ (10% mol cat., 50 °C) and Hey-Hawkins¹⁷ ($[\text{Mo}(\text{CO})_5(\text{PH}_2\text{Fc})]$, 10% mol cat., 66 °C), respectively. Moreover, Waterman's zirconium catalyst,^{4g} the ruthenium complex from Rosenberg,¹⁸ and the nickel complex from Webster¹⁹ required around 18–24 h with catalyst loadings of 5–10%. Furthermore, no byproducts derived from the insertion of more than one molecule of alkene (telomerization) were observed in our case.^{8d}

Similar reaction times were observed if an EWG is incorporated at the β -position of the olefin, as observed for dimethyl fumarate as a substrate (fum, entries 6 and 7). A small influence of the *cis/trans* geometry on the reaction-rates was observed, since slightly longer reaction times were observed for dimethylmaleate (the *cis*-isomer of dimethyl fumarate, entry 8). Lower catalytic loadings (e.g., 1% mol cat.) with these substrates were associated with lateral reactions that destroyed the catalyst. Nonetheless, complex **1** was robust enough to allow the reutilization of the catalytic solutions (5% mol cat.). Thus, after completion of the first catalytic run, dimethyl fumarate and PPh_2 were added again for five consecutive cycles without a significant decrease in the activity (see the Supporting Information).

Incorporation of a methyl group at the α -position (entry 9) decreases the reaction rate considerably, which can be attributed to both the increase of the steric effects and the decrease of the positive charge on the α -carbon.

The reaction is also tolerant to aldehydes, a poorly studied substrate in hydrophosphanation.²⁰ The corresponding phosphanes were obtained in less than 1 h at rt. The catalysis with cinnamaldehyde (entry 10) with a more EWG than crotonaldehyde (entry 11) was found to be faster, suggesting that electronic effects are more important than steric ones. Small amounts (less than 1% mol) of the corresponding 1-hydroxy-1,3-diphosphanes derived from the hydrophosphanation of the $\text{C}=\text{O}$ group were observed. Moreover, the diphosphane from cinnamaldehyde was almost quantitatively formed (91.4% yield) in 55 min if the catalysis was carried out with a 1:10:40 ratio (1:aldehyde: PPh_2), while 73.6% of the diphosphane was formed when using crotonaldehyde under similar conditions even after 18 h of reaction.

Less-activated substrates with a ketone group such as benzylideneacetone (entry 12) and cyclohexenone (entry 13) required longer reaction times, whereas no reaction was observed with nonactivated substrates, such as styrene or 2,3-

dimethyl-1-butene, even warming at 80 °C. It is interesting to mention that enones as substrates have been typically hydrophosphanated by palladium complexes as catalyst,^{4j,21} whereas a pincer-cobalt complex has been recently reported as suitable catalyst for the hydrophosphanation of the underexplored cyclohexenone.^{4e}

From the catalytic solutions, the functionalized phosphanes were separated by chromatography and isolated in almost quantitative yields (see the Supporting Information).

DFT Studies. The extraordinary activity of complex **1** in the hydrophosphanation of methylacrylate (MeAC, entry 2) prompted us to analyze the first step of the catalytic cycle by DFT-methods, i.e., the reaction between **1** and MeAC, whose energy profile is shown in Figure 1.

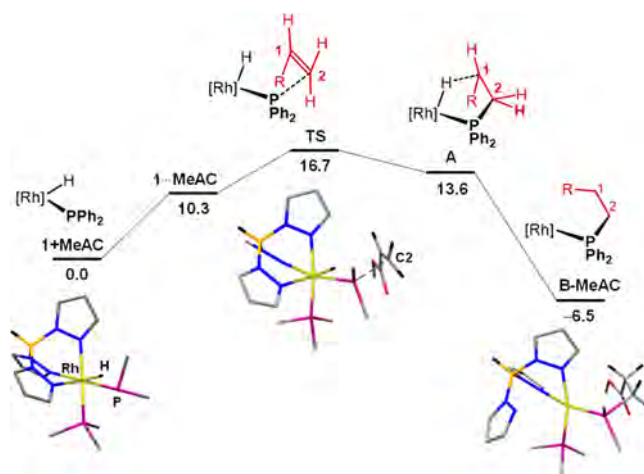


Figure 1. DFT computed (B3LYP-D3, 6-311G(d,p)/LanL2TZ(f)) Gibbs energy profile for the reaction of $[\text{Rh}(\text{Tp})\text{H}(\text{PMe}_3)(\text{PPh}_2)]$ (**1**) with methylacrylate. Relative ΔG_{298} values are given in kcal mol^{-1} . Only the C^{ipso} of the diphenylphosphanido ligand is shown for clarity. Selected bond distances (Å) are the following: P–C2 3.567 (**1**...MeAC), 2.387 (TS), 1.949 (A), 1.862 (B-MeAC); H–C1 2.617 (**1**...MeAC), 2.601 (TS), 2.283 (A). Color code is the following: Rh (yellow), O (red), N (blue), C (gray), B (orange), H (black). $[\text{Rh}] = [\text{Rh}(\text{Tp})(\text{PMe}_3)]$, R = CO_2Me .

After its formation, the initial adduct **1**...MeAC evolves to the intermediate **A** through a very accessible transition state **TS** (placed 16.7 kcal mol^{-1} above **1** + MeAC) featuring an incipient P–C bond. Intermediate **A** shows an almost strictly planar five-membered phosphametallacycle $\text{Rh}-\text{P}-\text{C}2-\text{C}1-\text{H}$.²² In addition, the sum of the three bond angles around C1 and C2 is $\sum^\circ = 360.0^\circ$ and 333.0° , respectively. Therefore, C2 has an almost tetrahedral geometry (with the new P–C bond), whereas C1 retains the sp^2 hybridization. Moreover, the remaining p orbital of C1 directly points at the hydrido ligand, an indication of interaction between C1 and the hydride, clearly observed in the HOMO of the complex **A** (Figure 2). This additional $\text{Rh}-\text{H}\cdots\text{C}1$ interaction should stabilize intermediate **A**, providing thus an easy path for the P–C bond formation step. From **A**, the hydride transfer from rhodium to C1 takes place directly. This transfer gives the bis(phosphane)rhodium(I) intermediate, $[\text{Rh}(\text{Tp})(\text{PMe}_3)(\text{PRPh}_2)]$ (R = $\text{CH}_2\text{CH}_2\text{CO}_2\text{Me}$, B-MeAC), which was found to be $-6.5 \text{ kcal mol}^{-1}$ more stable than **1** + MeAC.

The energy profile for this first step in the case of dimethyl fumarate (fum) was found to be quite similar, although in this

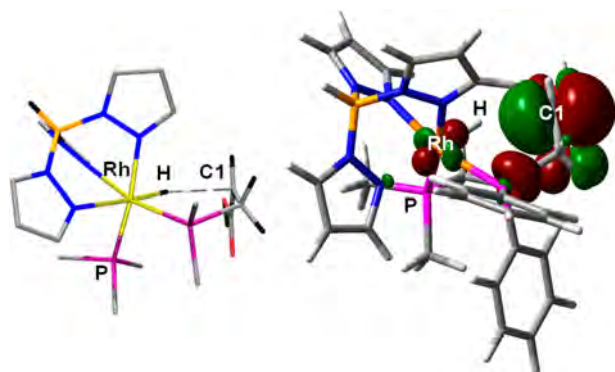


Figure 2. DFT computed geometry (B3LYP-D3, 6-311G(d,p)/LanL2TZ(f)) for intermediate A (left) and HOMO of the complex (right).

case species A is not an intermediate but a transition state TS' (Figure 3) placed 19.9 kcal mol⁻¹ above 1 + fum.

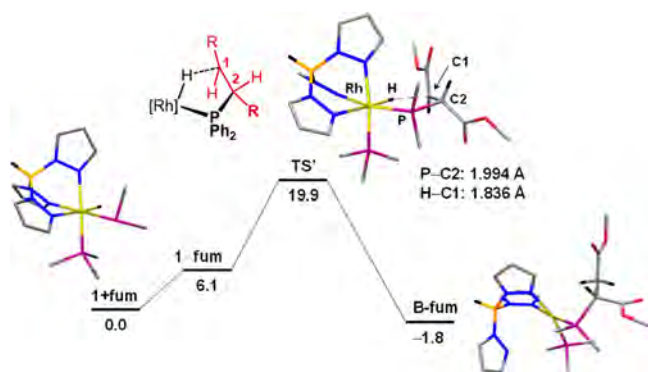


Figure 3. DFT computed (B3LYP-D3, 6-311G(d,p)/LanL2TZ(f)) Gibbs energy profile for the reaction of [Rh(Tp)H(PMe₃)(PPh₂)] (1) with dimethyl fumarate (fum). Relative $\Delta G_{298}^{\ddagger}$ values are given in kcal mol⁻¹. Only the C^{IPSO} of the diphenylphosphanido ligand is shown for clarity. Color code is the following: Rh (yellow), O (red), N (blue), C (gray), B (orange), H (black). [Rh] = "Rh(Tp)(PMe₃)", R = CO₂Me.

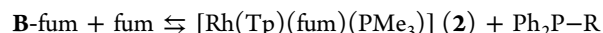
Noticeably, TS' results from a concerted approach of the two olefinic carbons of dimethyl fumarate to the hydride and the phosphanido groups to give a five-membered phosphametallacycle similar to that found in the intermediate A (Figure 1). Through TS', the adduct 1•fum directly transforms into the rhodium(I) bis(phosphane) intermediate [Rh(Tp)(PMe₃)(PRPh₂)] (R = CH(CO₂Me)CH₂CO₂Me, B-fum, Figure 3) in a concerted outer-sphere single step.

The findings disclosed by DFT studies are remarkable. First and to the best of our knowledge, we are not aware of previous proposals of outer-sphere single-step mechanisms for the critical P–C bond formation, and second, the relevant role of the hydride ligand is unprecedented in olefin hydrophosphanation. Moreover, the beneficial Rh–H···C1 interaction observed in intermediate A (Figure 2) and TS' (Figure 3) requires an electrophilic hydride because of the negative charge accumulated on C1. In this scenario it is expected that acrylonitrile (with a more EWG at the α -position than methylacrylate)¹⁵ is hydrophosphanated more slowly than MeAC (as experimentally observed), as a result of the partial reduction of the negative charge at the α -carbon that weakens the positive Rh–H···C1 interaction. On the whole, the P–Rh–

H moiety seems to act as a 1,3-dipole in these reactions, which could be considered as 1,3-dipolar cycloadditions. Furthermore, addition of ^tBuOH (5 mmol) to the catalytic solutions in the acrylonitrile and dimethyl fumarate cases did not produce appreciable changes in the reaction rate, supporting again the "click" scenario (1,3-cycloaddition) for the key P–C and C–H bond formation steps. Notice that if a free carbanion were present, the catalysis should be accelerated by ^tBuOH.^{8d}

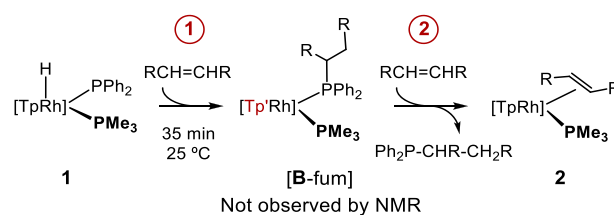
Stoichiometric Studies. Further information about the reliability of the DFT profiles in Figures 1 and 3 as well as the fundamental role of the hydride ligand was gathered through several additional selected experiments at stoichiometric level.

In the case of dimethyl fumarate, for which the equilibrium



is fully shifted to the right ($\Delta G^{\circ} = -3.1$ kcal mol⁻¹ from DFT), the reaction of [Rh(Tp)H(PMe₃)(PPh₂)] (1) with dimethyl fumarate was analyzed by NMR spectroscopy (Scheme 2). It

Scheme 2. Reaction of [Rh(Tp)H(PMe₃)(PPh₂)] (1) with Dimethyl Fumarate, R = CO₂Me, Tp' = κ^2 -Tp



required 2 mol-equiv of the olefin to reach completion, and the products were found to be [Rh(Tp)(fum)(PMe₃)] (2)²³ and the functionalized phosphane, Ph₂P–R (R = CH(CO₂Me)CH₂CO₂Me).

Interestingly, if the reaction is carried out in a 1:1 molar ratio, complex 2 and the free phosphane were again obtained, but in this case 50% of complex 1 remains unreacted. This observation provides a clear indication that step 1 is the rate-determining step. Then, the reaction rate can be expressed as²⁴

$$\text{rate} = -\frac{d[\mathbf{1}]}{dt} = k_1[\mathbf{1}][\text{fum}] = 2k_1[\mathbf{1}]^2$$

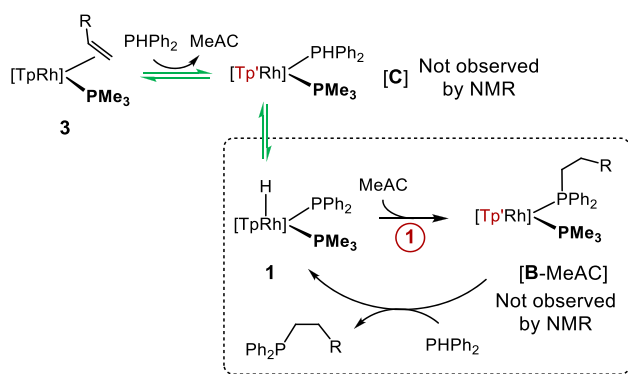
A plot of 1/[1] vs time proved to be linear with a slope = 2k₁ (see the Supporting Information). The value of k₁ = 4.65 × 10⁻² M⁻¹ s⁻¹ translates to a free energy barrier of ca. 19.3 kcal mol⁻¹ (298 K) using the Eyring–Polanyi equation.²⁵

Goodness of this single-point calculation for ΔG^{\ddagger} was experimentally confirmed by carrying out a full study of the reaction in the temperature range 213–307 K (Eyring plot). The activation parameters $\Delta H^{\ddagger} = 6.3 \pm 0.2$ kcal mol⁻¹ and $\Delta S^{\ddagger} = -44.3 \pm 0.6$ cal mol⁻¹ K⁻¹ (extracted from the straight line obtained) gave a value for $\Delta G_{298}^{\ddagger} = 19.5$ kcal mol⁻¹, which is similar to that estimated from the single-point approximation. Both values for ΔG^{\ddagger} are in excellent agreement with that obtained from DFT calculations (19.9 kcal mol⁻¹). Moreover, the large and negative activation entropy can be seen as diagnostic of a highly ordered transition state as that shown in Figure 3. Furthermore, a parallel experiment with [Rh(Tp)D(PMe₃)(PPh₂)] (1^D) and dimethyl fumarate (in 1:2 molar ratio) under identical conditions gave a ratio k_H/k_D = 1.31. This small kinetic isotope effect (KIE) agrees with the simple elongation of the Rh–H bond in the transition state

TS' (Figure 3). Accordingly, a value of 1.3 was estimated from DFT-studies.

For methyl acrylate (MeAC), the corresponding equilibrium $\text{B-MeAC} + \text{MeAC} \rightleftharpoons [\text{Rh}(\text{Tp})(\text{MeAC})(\text{PMe}_3)]$ (**3**) + $\text{Ph}_2\text{P-R}'$ ($\text{R}' = \text{CH}_2\text{CH}_2\text{CO}_2\text{Me}$) is fully shifted to the left ($\Delta G^0 = +2.7 \text{ kcal mol}^{-1}$ from DFT) so that attention was paid on the reaction of $[\text{Rh}(\text{Tp})(\text{MeAC})(\text{PMe}_3)]$ (**3**) with diphenylphosphane. This reaction also requires 2 mol-equiv of PPh_2 to reach completion, and the products were found to be $[\text{Rh}(\text{Tp})\text{H}(\text{PMe}_3)(\text{PPh}_2)]$ (**1**) and $\text{Ph}_2\text{P-R}'$ (Scheme 3). In

Scheme 3. Reaction of $[\text{Rh}(\text{Tp})(\text{MeAC})(\text{PMe}_3)]$ (**3**) with PPh_2 , $\text{R} = \text{CO}_2\text{Me}$, $\text{Tp}' = \kappa^2\text{-Tp}$



this case, replacement of MeAC by PPh_2 to render intermediate $[\text{Rh}(\text{Tp})(\text{PPh}_2)(\text{PMe}_3)]$ (**C**) and then $[\text{Rh}(\text{Tp})\text{H}(\text{PMe}_3)(\text{PPh}_2)]$ (**1**) was so fast that complex **3** was not present anymore in the first ^1H NMR spectra (recorded after 3.5 min of reaction). Moreover, in this spectrum complex **1** and the phosphane $\text{Ph}_2\text{P-R}'$, along with small amounts of free methyl acrylate and PPh_2 in a stoichiometric ratio, were present (Scheme 3).

Because the reaction reached completion after ca. 15 min, four ^1H NMR spectra were recorded; evolution of the concentration of free methyl acrylate vs time was found to be exponential, whereas the concentration of **1** remained almost constant (see the Supporting Information).

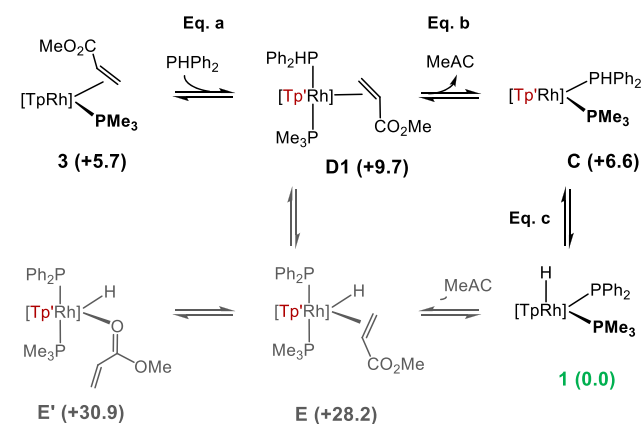
Consequently, what we were indeed observing in the monitored 15 min corresponds to the catalytic formation of $\text{Ph}_2\text{P-R}'$ by complex **1** in a scenario where $[\mathbf{1}] \approx [\mathbf{3}]^0$ (Scheme 3). Therefore,

$$\text{rate} = k_1[\mathbf{1}][\text{MeAC}] = k_1[\mathbf{3}]^0[\text{MeAC}] = k_{\text{obs}}[\text{MeAC}]$$

The value of $k_{\text{obs}} = 6.71 \times 10^{-3} \text{ s}^{-1}$ corresponds to $k_1 = 1.18 \text{ M}^{-1} \text{ s}^{-1}$, which translates to a free energy barrier of ca. $17.3 \text{ kcal mol}^{-1}$ (298 K) for step 1 in Scheme 3 by using the Eyring–Polanyi equation. This value fits considerably well with that obtained from DFT studies ($16.7 \text{ kcal mol}^{-1}$ (298 K), Figure 1). Moreover, since the full transformation of $[\text{Rh}(\text{Tp})(\text{MeAC})(\text{PMe}_3)]$ (**3**) into $[\text{Rh}(\text{Tp})\text{H}(\text{PMe}_3)(\text{PPh}_2)]$ (**1**) was found to be faster than the P–C bond formation reaction (Scheme 3), it can be concluded that the activation energy for the P–H bond cleavage is lower than $17.3 \text{ kcal mol}^{-1}$, and consequently, the previously reported value of ca. 20 kcal mol^{-1} (from DFT) was most probably overestimated.¹³

Additionally, the first part of the reaction (from **3** to **1**) was analyzed by DFT calculations. The first step (eq a, Scheme 4) requires dissociation of a pyrazolate arm in **3** to allow the coordination of diphenylphosphane, which renders accessible

Scheme 4. Elemental Steps for the Transformation of **3** into **1** and Plausible Intermediates for Alternative Inner-Sphere Mechanisms, $\text{Tp}' = \kappa^2\text{-Tp}$ ^a



^aThermodynamic Gibbs energy values (DFT computed, B3LYP-D3, 6-311G(d,p)/LanL2TZ(f)) relative to **1** + MeAC (in green) are shown in parentheses, in kcal mol^{-1} .

intermediates of the type $[\text{Rh}(\text{Tp}')(\text{PMe}_3)(\text{PPh}_2)(\text{MeAC})]$ (**D**, $\text{Tp}' = \kappa^2\text{-Tp}$); the *trans* isomer (**D1**, Scheme 4) was found to be slightly more stable than the *cis* counterpart in $0.9 \text{ kcal mol}^{-1}$. From **D**, extrusion of the olefin gives the square-planar intermediate $[\text{Rh}(\text{Tp}')(\text{PMe}_3)(\text{PPh}_2)]$ (**C**), in which an easy P–H bond activation reaction produces complex **1**. This step (eq c) is the driving force for the full transformation of **3** into **1**.

At this point, the alternative inner-sphere mechanism was tested by modeling the “hypothetical” intermediate $[\text{Rh}(\text{Tp}')\text{-H}(\text{PMe}_3)(\text{PPh}_2)(\text{olefin})]$ (**E**), which would be the key species for further hydride insertion and P–C bond formation reactions (Scheme 1). Complex **E** could be the result of the P–H bond cleavage in **D**, or alternatively, it can arise from the dissociation of a pyrazolate arm in **1** to allow olefin coordination. In any case, complex **E** was found to be $28.2 \text{ kcal mol}^{-1}$ above **1** + MeAC, and consequently, this alternative was discarded. Moreover, the “hypothetical” isomer with methyl acrylate O-coordinated (**E'**) was also modeled since this complex would be the first step for a further intramolecular nucleophilic attack of the phosphanido to the olefin, as proposed by Leung in palladium chemistry.^{21a,26} However, complex **E'** was found to be even higher in energy than **E** (Scheme 4). Moreover, both complexes (**E** and **E'**) lie above the transition state ($16.7 \text{ kcal mol}^{-1}$) shown in Figure 1 for the novel “click step” reported here, which is the lowest energy pathway we have found by DFT calculations.

The last point analyzed at stoichiometric level concerns the prominent role of the hydride ligand commented above. To this end, attention was focused on the related complex $[\text{Rh}(\text{Tp})\text{Cl}(\text{PMe}_3)(\text{PPh}_2)]$ (**4**), in which the hydride ligand in **1** has been formally replaced by a chloride one.

Complex **4** was prepared from the reaction of $[\text{Rh}(\text{Tp})\text{H}(\text{PMe}_3)(\text{PPh}_2)]$ (**1**) with $\text{ClCH}_2\text{CO}_2\text{Me}$ (see the Supporting Information) and further characterized by analytical and spectroscopic data, including an X-ray diffraction analysis (Figure 4). In the complex, the rhodium atom shows a distorted octahedral environment bound to a *fac*-Tp ligand and to trimethylphosphane, the terminal phosphanido, and chlorido ligands. The geometry around the phosphanido

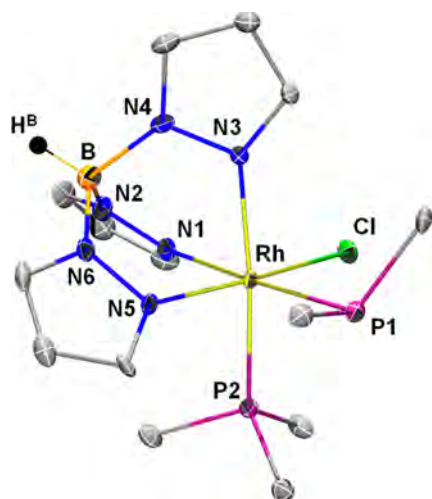


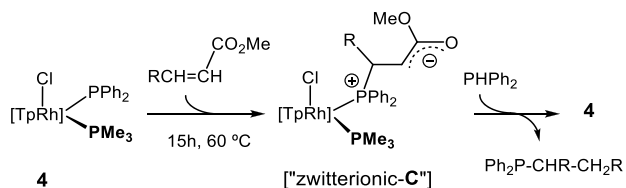
Figure 4. Molecular structure (ORTEP, ellipsoids set at 50% probability) of complex $[\text{Rh}(\text{Tp})\text{Cl}(\text{PMe}_3)(\text{PPh}_2)]$ (**4**). One of the two similar independent molecules is shown. Selected bond distances [Å] and angles [deg] are the following: Rh–P1 2.404(2)/2.399(2), Rh–P2 2.283(2)/2.288(2), Rh–Cl 2.356(2)/2.357(2), Rh–N1 2.180(6)/2.191(7), Rh–N3 2.131(6)/2.131(7), Rh–N5 2.075(7)/2.073(8), P1–Rh–N1 176.7(2)/175.0(2), P2–Rh–N3 174.2(2)/173.9(2), Cl–Rh–N5 174.4(2)/175.0(2). Only the C^{ipso} carbons of the phenyl groups are shown for clarity.

phosphorus atom (P1) reveals that the lone pair does not interact with Rh resulting in a Rh–P1 single bond. Thus, the environment of P1 is pyramidal with the sum of the three bond angles amounting to $\sum^\circ = 314.6(3)/316.2(3)^\circ$.

According to its formulation, complex **4** shows two signals corresponding to the PPh_2 and PMe_3 groups as doublets of doublets in the $^{31}\text{P}\{^1\text{H}\}$ NMR spectrum at $\delta = 56.3$ ($J(\text{P},\text{Rh}) = 61$ Hz) and 0.2 ($J(\text{P},\text{Rh}) = 119$ Hz) ppm, respectively. The value of $J(\text{P},\text{Rh})$ for the terminal phosphanido ligand was found to be smaller than those corresponding to the phosphane. This difference can be attributed to a substantial reduction in the σ -orbital character of the Rh– PR_2 bond as compared to the Rh–phosphane,¹² providing thus a useful tool for the identification of the terminal phosphanido ligand.

The reaction of $[\text{Rh}(\text{Tp})\text{Cl}(\text{PMe}_3)(\text{PPh}_2)]$ (**4**) with dimethyl fumarate was carried out in the presence of 1 mol of PPh_2 to attain protonolysis of the expected carbanion that regenerates complex **4** (Scheme 5).

Scheme 5. Reaction of $[\text{Rh}(\text{Tp})\text{Cl}(\text{PMe}_3)(\text{PPh}_2)]$ (**4**) with Dimethyl Fumarate and Diphenylphosphane, $\text{R} = \text{CO}_2\text{Me}$

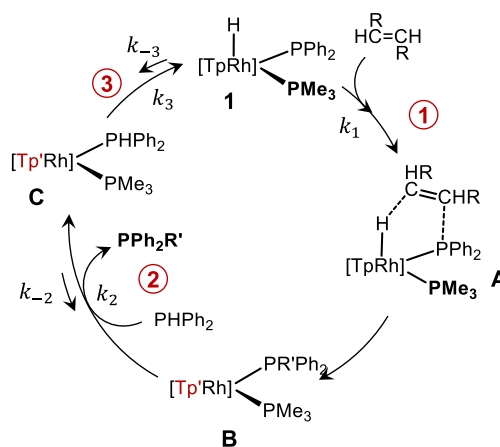


The reaction succeeded, but it required around 15 h at 60 °C to reach completion.²⁷ This experiment definitively confirms the beneficial role of the hydrido ligand in the facile hydrophosphanation of dimethyl fumarate with the hydrido complex **1**.

Kinetic Studies and Proposed Mechanism. On the basis of these experimental and theoretical results, we propose

the catalytic cycle shown in Scheme 6. The first step corresponds to the interaction of catalyst **1** with the olefin to

Scheme 6. Proposed Catalytic Cycle for Olefin Hydrophosphanation with PPh_2 Catalyzed by Complex **1**, $\text{Tp}' = \kappa^2\text{-Tp}$



yield intermediate **A** (TS' for dimethyl fumarate), which directly evolves to the bis(phosphane)rhodium(I) complex **B**. Next, a ligand exchange reaction of the functionalized phosphane by PPh_2 gives intermediate **C**, for which an easy P–H bond activation reaction (almost immediate at rt¹²) regenerates the catalyst, closing thus the catalytic cycle.

Further support for the proposed catalytic cycle was gathered from kinetic studies. A first analysis of the catalysis under standard conditions [cat.]:[olefin]:[PPh_2] 1:20:20 (1:5000:5000 for MeAC) revealed mainly two different behaviors. For substrates whose entry is marked in red (Table 1), plots of [olefin] vs time gave very good exponential fittings, whereas for those marked in green, an excellent fit to straight lines were obtained when plotting $[\text{PH}]^0\text{Ln}[\text{PH}] - [\text{PH}]$ ($[\text{PH}] = [\text{PPh}_2]$) vs time. Figure 5 shows a couple of representative examples, while all of them can be found in the Supporting Information.

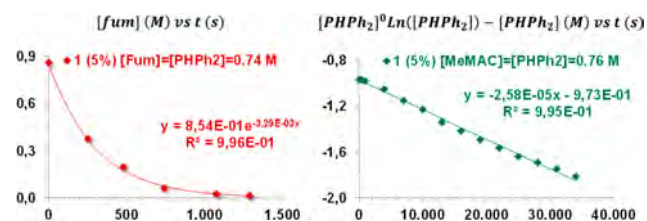


Figure 5. Plots of [fum] (fum = dimethyl fumarate, left) vs time (s) and $[\text{PH}]^0\text{Ln}[\text{PH}] - [\text{PH}]$ ($[\text{PH}] = [\text{PPh}_2]$, right) vs time (s) for the catalytic hydrophosphanation of dimethyl fumarate (in red) and methyl methacrylate (in green) using complex **1** as catalyst.

Moreover, a more detailed study on the hydrophosphanation of dimethyl fumarate and acrylonitrile (both marked in red in Table 1) revealed some interesting differences. Thus, for dimethyl fumarate, a zeroth-order dependence of the reaction rate on $[\text{Ph}_2\text{PH}]$ and first-order dependence on [fum] and on **1** were obtained, whereas in the acrylonitrile case, a first-order dependence was observed for the three components $[\text{PPh}_2]$, $[\text{ACN}]$, and **1** (see the Supporting Information).

Accordingly, plots of $\ln(V^0/[I]^0)$ versus $\ln([\text{P}(\text{HPh}_2)]^0)$, $\ln(V^0/[I]^0)$ versus $\ln([\text{fum}]^0)$, and $\ln(k_{\text{obs}})$ versus $\ln([\text{I}]^0)$ were linear with slopes of 0.083, 0.92, and 0.99, respectively, for the hydrophosphonation of dimethyl fumarate, while these values transform into 1.15, 1.16 and 1.17, respectively, for acrylonitrile hydrophosphonation.

This apparent contradiction of different rate-laws for the same catalytic cycle can be rationalized according to the rate law corresponding to the catalytic cycle shown in Scheme 6:

$$\text{rate} = -\frac{d[\text{ole}]}{dt} = \frac{k_1 k_2 [\text{I}]^0 [\text{PH}][\text{ole}]}{k_1 [\text{ole}] + k_2 [\text{PH}] + \frac{k_{-2} k_1}{k_3} [\text{ole}][\text{PR}]}$$

where ole = olefin, PH = $\text{P}(\text{HPh}_2)$, PR = functionalized phosphane, from which three main scenarios have been observed.

(i) The first possibility involves the term $k_2[\text{PH}]$ in the denominator being larger than the other two in such a way that the rate law simplifies to $\text{rate} = k_1[\text{I}]^0[\text{ole}]$, and it can be expected for substrates that benefit from a rapid phosphane exchange reaction (step 2) and relatively slow step 1. Since this exchange should be easier with less electron-rich functionalized phosphanes, it is not strange that dimethyl fumarate and dimethyl maleate, both leading to $\text{Ph}_2\text{P}-\text{CH}(\text{CO}_2\text{Me})-\text{CH}_2\text{CO}_2\text{Me}$ (the most electron-poor phosphane of those considered), follow this rate law. Accordingly, a zeroth-order dependence of the reaction rate on $[\text{Ph}_2\text{PH}]$ and first-order dependence on $[\text{olefin}]$ and on $[\text{I}]$ are expected, as experimentally observed in the case of dimethyl fumarate. For these substrates, step 1 is thought to be the rate-determining step in full agreement with the high value of the experimentally measured entropy (see above).

(ii) In the case of substrates for which

$$k_2[\text{PH}] + k_1[\text{ole}] \gg \frac{k_{-2} k_1}{k_3} [\text{ole}][\text{PR}] \quad (k_1 \approx k_2)$$

the rate law simplifies under catalytic conditions ($[\text{PH}]^0 = [\text{ole}]^0 = [\text{S}]^0$) to

$$\text{rate} = \frac{k_1 k_2 [\text{I}]^0 [\text{PH}][\text{ole}]}{k_1 [\text{ole}] + k_2 [\text{PH}]} = \frac{k_1 k_2}{k_1 + k_2} [\text{I}]^0 [\text{S}] = k_{\text{cat}} [\text{I}]^0 [\text{S}]$$

Consequently, plots of $[\text{olefin}]$ vs time are expected to give exponential fittings under catalytic conditions. Moreover, in the presence of an excess of olefin,

$$k_1[\text{ole}]^0 > k_2[\text{PH}] \Rightarrow \text{rate} = k_2[\text{I}]^0[\text{PH}]$$

whereas in the presence of an excess of $\text{P}(\text{HPh}_2)$,

$$k_2[\text{PH}]^0 > k_1[\text{ole}] \Rightarrow \text{rate} = k_1[\text{I}]^0[\text{ole}]$$

so that a first order dependence on both olefin and $\text{P}(\text{HPh}_2)$ is expected under pseudo-first-order conditions. This scenario corresponds to that experimentally found for acrylonitrile (and most probably for methyl acrylate). As a matter of fact, Figure 6 shows a couple of catalysis under identical pseudo-first-order conditions for $\text{P}(\text{HPh}_2)$ (in black) and acrylonitrile (in red).

A nice exponential fitting was obtained for both experiments, with almost identical values of $k_{\text{obs}} = k_1[\text{I}]^0$ and $k_2[\text{I}]^0$, which clearly support the initial premise of $k_1 \approx k_2$.

(iii) For substrates that show the reverse situation with equilibrium 2 shifted to species B, for which

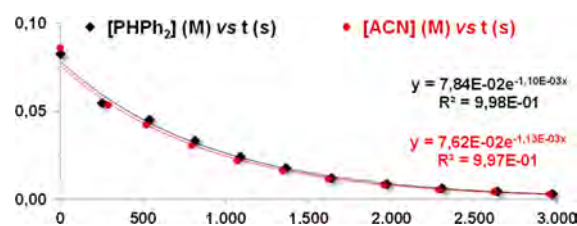


Figure 6. Plots of $[\text{P}(\text{HPh}_2)]$ (M) (in black) and $[\text{ACN}]$ (M) (in red) vs time (s) under pseudo-first-order conditions: $[\text{I}]^0$: $[\text{ACN}]^0$: $[\text{P}(\text{HPh}_2)]^0 = 1:50:10$ in black and $1:10:50$ in red. $[\text{I}]^0 = 0.0083$ M in both cases, taken from a stock solution.

$$\frac{k_{-2} k_1}{k_3} [\text{ole}][\text{PR}] \gg k_2[\text{PH}] + k_1[\text{ole}]$$

the rate law becomes

$$\text{rate} = k_3 K_{\text{eq}2} [\text{I}]^0 \frac{[\text{PH}]}{[\text{PH}]^0 - [\text{PH}]}$$

and it is followed by substrates marked in green in Table 1 ($k_{\text{obs}} = k_3 K_{\text{eq}2} [\text{I}]^0$), which hydrophosphanated more slowly by catalyst **1** than the red ones due to the poisoning of the catalyst by the product to some extent. Indeed, the resulting phosphanes in these cases are electronically richer than those arising from methyl acrylate, acrylonitrile, and dimethyl fumarate/maleate.

Additionally, the progress of the selected catalysis in the presence of an excess of equimolar amounts of $\text{P}(\text{HPh}_2)$ and $\text{P}(\text{DPh}_2)$ (ratio cat.:fum: $\text{P}(\text{HPh}_2)$: $\text{P}(\text{DPh}_2) = 1:20:100:100$, intermolecular competition²⁸) was monitored to determine the isotope effect. This method was chosen mainly because olefin functionalization by $\text{P}(\text{HPh}_2)$ and $\text{P}(\text{DPh}_2)$ occurs under exactly the same conditions, avoiding experimental errors, and because the ratio of the functionalized phosphanes ($\text{PR}_{\text{H}}/\text{PR}_{\text{D}}$) can be measured by NMR spectroscopy with very good precision. Values of $\text{PR}_{\text{H}}/\text{PR}_{\text{D}} = 1.28, 1.22, \text{ and } 1.27$ were respectively observed for methyl acrylate, acrylonitrile, and dimethyl fumarate (three substrates marked in red in Table 1, scenarios i and ii), which agree with that obtained for step 1 (dimethyl fumarate) at stoichiometric level. On the contrary, a larger value of 2.33 was measured for cyclohexanone (marked in green, scenario iii), which could be related to the cleavage of the P–H bond (or alternatively, previous $\text{P}(\text{HPh}_2)$ coordination to intermediate B),²⁸ in the rate-determining step.

Interestingly, in the particular case of dimethyl fumarate and cyclohexanone, where two chiral carbons are formed in the hydrophosphonation process, one enantiomeric pair was formed in our case, as evidenced by ^1H and ^2H NMR spectroscopy (see the Supporting Information). Most probably, the observed enantiomeric pair corresponds to that arising from the *syn*-addition of the hydrido and phosphanido ligands to the olefin, as expected from the click-step shown in Figures 1 and 3. Note that in a stepwise scenario, through a zwitterionic intermediate with a carbanion as proposed by Glueck,^{3a} free rotation around the C1–C2 bond would allow the entry of the hydrido ligand on the other side so that the second enantiomeric pair (corresponding to the *anti*-addition) can be expected, at least in some extension. In other words, the observation of a single enantiomeric pair corroborates the outer-sphere single-step mechanism for the critical P–C bond formation reported here.

Finally, the plausible inhibition of the catalyst by the olefin, suitable to sequester part of the catalyst as the rhodium(I) complexes $[\text{Rh}(\text{Tp})(\text{olefin})(\text{PMe}_3)]$, was tested estimating (by DFT studies) the thermodynamic free energies for the equilibria: $[\text{Rh}(\text{Tp})(\text{olefin})(\text{PMe}_3)] + \text{PPh}_2 \rightleftharpoons \mathbf{1} + \text{olefin}$.

Values of $\Delta G^\circ = -1.2$, -5.7 , and -4.8 were obtained for dimethyl fumarate, methyl acrylate, and acrylonitrile, respectively. Consequently, the equilibria are shifted to right and the main catalytic cycle that operates is that shown in Scheme 6.

CONCLUSIONS

The hydrido-phosphanido rhodium complex $[\text{Rh}(\text{Tp})\text{H}(\text{PMe}_3)(\text{PPh}_2)]$ (**1**), containing a tripodal ligand (Tp) and an additional phosphane ligand, readily catalyzes the addition of diphenylphosphane to a wide range of olefins. Its good activity likely arises from a catalytic cycle including three well-separated steps: (i) an outer-sphere P–C and H–C bond formation likely through a 1,3-dipolar cycloaddition, (ii) a replacement reaction of the functionalized phosphane by PPh_2 , and (iii) an easy P–H bond activation reaction that regenerates the catalyst. Stoichiometric, kinetic, and DFT studies provided further support for this picture, in which the critical P–C bond formation step takes place through intermediates (or TS) in which both the phosphanido and the hydrido ligands interact with the α - and β -carbons of the alkene, respectively, in a cooperative way. The beneficial role of the Rh–H...C1 interaction, which avoids side-reactions (as telomerization) and facilitates the hydrogen transfer from rhodium to C1, is remarkable. In full agreement, the related complex $[\text{Rh}(\text{Tp})\text{Cl}(\text{PMe}_3)(\text{PPh}_2)]$ (**4**), which lacks the electrophilic arm (the hydride present in **1**), was found to be a poor catalyst. We believe that the findings reported here have important implications for hydrophosphanation reactions, especially in scenarios associated with P–H bond cleavage steps, and represent substantial progress from a mechanistic point of view. Knowing that both the nucleophilicity of the phosphanido ligand and the electrophilicity²⁹ of the hydrido ligand play key roles may help develop new catalysts for the “green syntheses” of phosphanes and related reactions. Further studies on this topic are currently in progress.

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/jacs.0c11010>.

Experimental details, catalytic and kinetic studies, and selected NMR spectra of complexes and phosphanes (PDF)

Cartesian coordinates for optimized (DFT) structures of complexes, intermediates, and transition states (XYZ)

Crystallographic data for complex **4** (CIF)

Crystallographic data for complex $[\text{Rh}(\text{Tp})\text{Cl}_2(\text{PMe}_3)]$ (CIF)

AUTHOR INFORMATION

Corresponding Author

Cristina Tejel – Departamento de Química Inorgánica, Instituto de Síntesis Química y Catálisis Homogénea (ISQCH), CSIC-Universidad de Zaragoza, 50009 Zaragoza, Spain; orcid.org/0000-0003-3306-0635; Email: ctejel@unizar.es

Authors

Víctor Varela-Izquierdo – Departamento de Química Inorgánica, Instituto de Síntesis Química y Catálisis Homogénea (ISQCH), CSIC-Universidad de Zaragoza, 50009 Zaragoza, Spain; orcid.org/0000-0002-7806-8959

Ana M. Geer – Departamento de Química Inorgánica, Instituto de Síntesis Química y Catálisis Homogénea (ISQCH), CSIC-Universidad de Zaragoza, 50009 Zaragoza, Spain; orcid.org/0000-0003-1115-6759

Janeth Navarro – Departamento de Química Inorgánica, Instituto de Síntesis Química y Catálisis Homogénea (ISQCH), CSIC-Universidad de Zaragoza, 50009 Zaragoza, Spain; orcid.org/0000-0002-2549-4487

José A. López – Departamento de Química Inorgánica, Instituto de Síntesis Química y Catálisis Homogénea (ISQCH), CSIC-Universidad de Zaragoza, 50009 Zaragoza, Spain; orcid.org/0000-0003-4500-1974

Miguel A. Ciriano – Departamento de Química Inorgánica, Instituto de Síntesis Química y Catálisis Homogénea (ISQCH), CSIC-Universidad de Zaragoza, 50009 Zaragoza, Spain; orcid.org/0000-0002-7188-8014

Complete contact information is available at: <https://pubs.acs.org/10.1021/jacs.0c11010>

Author Contributions

[†]V.V.-I., A.M.G., and J.N. contributed equally.

Notes

The authors declare no competing financial interest.

The crystallographic data have been deposited with the Cambridge Crystallographic Data Centre (1965870 and 1965871).

ACKNOWLEDGMENTS

The generous financial support from AEI/FEDER, UE (CTQ2017-83421-P), Gobierno de Aragón/FEDER, UE (GA/FEDER, Inorganic Molecular Architecture Group E08_17R) is gratefully acknowledged. A.M.G. and V.V.-I. thank MINECO/FEDER for a JdC contract and a FPI fellowship, respectively. The Centro de Supercomputación de Galicia (CESGA) is also gratefully acknowledged for generous allocation of time.

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- (24) Because complex **B-fum** was not observed, it can be assumed that $[\text{1}]^0 \approx [\text{1}] + [\text{2}]$ and $[\text{fum}]^0 = [\text{fum}] + [\text{2}] + [\text{PR}]$ (PR = functionalized phosphane), and since PR is formed simultaneously with complex **2** (through step 2) $\Rightarrow [\text{2}] = [\text{PR}]$ at any time. Taking into account the experimental conditions $[\text{fum}]^0 = 2[\text{1}]^0$, a simple calculation leads to $[\text{fum}] = 2[\text{1}]$ at any time. Consequently, the rate law becomes $\text{rate} = 2k_1[\text{1}]^2$.
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