

Chemistry of a Nitrosyl Ligand $\kappa\eta$ -Bridging a Ditungsten Center: Rearrangement and N–O Bond Cleavage Reactions

M. Angeles Alvarez, M. Esther García, Daniel García-Vivó,* Ana M. Guerra, Miguel A. Ruiz,* and Larry R. Falvello

Cite This: *Inorg. Chem.* 2022, 61, 14929–14933

Read Online

ACCESS |



Metrics & More



Article Recommendations



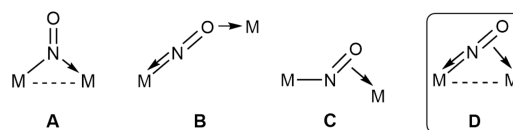
Supporting Information

ABSTRACT: The novel nitrosyl-bridged complex $[W_2Cp_2(\mu-P^tBu_2)(\mu-\kappa\eta\text{-NO})(CO)(NO)](BAR_4)$ [$Ar = 3,5\text{-}C_6H_3(CF_3)_2$] was prepared in a multistep procedure starting from the hydride $[W_2Cp_2(\mu-H)(\mu-P^tBu_2)(CO)_4]$ and involving the new complexes $[W_2Cp_2(\mu-P^tBu_2)(CO)_4](BF_4)$, $[W_2Cp_2(\mu-P^tBu_2)(CO)_2(NO)_2](BAR_4)$, and $[W_2(\mu-\kappa\eta^5\text{-}C_5H_4)Cp(\mu-P^tBu_2)(CO)(NO)_2]$ as intermediates, which follow from reactions with $HBf_4 \cdot OEt_2$, NO, and $Me_3NO \cdot 2H_2O$, respectively. The nitrosyl-bridged cation easily added chloride upon reaction with $[N(PPh_3)_2]Cl$, with concomitant NO rearrangement into the terminal coordination mode, to give $[W_2ClCp_2(\mu-P^tBu_2)(CO)(NO)_2]$, and underwent N–O and W–W bond cleavages upon the addition of CN^tBu to give the mononuclear phosphinoimido complex $[WCp(NP^tBu_2)(CN^tBu)_2](BAR_4)$. Another N–O bond cleavage was induced upon photochemical decarbonylation at 243 K, which gave the oxo- and phosphinito-bridged nitrido complex $[W_2Cp_2(N)(\mu-O)(\mu-OP^tBu_2)(NO)](BAR_4)$, likely resulting from a N–O bond cleavage step following decarbonylation.

Nitric oxide (NO) is a multifaceted molecule able to bind metal atoms in both high and low oxidation states with very diverse coordination modes, which makes the chemistry of nitrosyl complexes a research area of great academic interest.¹ This simple molecule also has relevant biological activities at low doses (neurotransmission, regulation of blood pressure, tumorigenic activity, etc.),² and some nitrosyl complexes can actually be designed as drugs releasing NO in a controlled way for therapeutic purposes.³ In contrast, the presence of NO in air at high levels has undesired consequences (toxicity, greenhouse effect, destruction of stratospheric ozone, etc.), which makes it an important air pollutant requiring catalytic, metal-mediated abatement.⁴ The latter often involves cleavage of the strong N–O bond of this molecule while interacting with one or more metal atoms at the surface of solid catalysts. Such elemental reactions thus become fundamental processes to be studied in search of potential improvements in the catalysis for NO abatement. This is why we have been studying for some time the chemistry of binuclear nitrosyl complexes bearing different types of unsaturation (coordinative, electronic, or both) because these are molecular systems potentially able to activate (weaken) and eventually cleave the N–O bond of nitric oxide at a dimetal site.^{5–7}

The nitrosyl ligand is known to bind two metal atoms in four distinct ways (A to D in Chart 1). In the more common N:N bridging (or semibridging) mode (A), the ligand provides the dimetal center with three electrons and is somewhat activated toward cleavage of its N–O bond, which is weakened with respect to the bond in terminal linear nitrosyls in mononuclear complexes. The coordination of type B involves additional binding of the oxygen atom to a second metal atom via lone electron pairs of oxygen and should not modify much the strength of the N–O interaction. In contrast, the bent (C) or

Chart 1. Coordination Modes of Nitrosyl Ligands in Binuclear Complexes



linear (D) $\kappa\eta$ modes involve coordination to the second metal atom via a $\pi(N-O)$ bonding orbital, and this would be expected to significantly weaken that bond. Unfortunately, only a couple of examples of these coordination modes are currently known,^{8,9} indeed displaying elongated N–O bonds of ca. 1.30 Å, and their chemistry remains unexplored. As a part of our studies on unsaturated binuclear nitrosyl complexes, we here report an efficient synthetic route for a new complex of type D, the cationic complex $[W_2Cp_2(\mu-P^tBu_2)(\mu-\kappa\eta\text{-NO})(CO)(NO)](BAR_4)$ ($Ar = 3,5\text{-}C_6H_3(CF_3)_2$) and a preliminary study of its chemical behavior. Interestingly, several N–O bond cleavage reactions can be induced at the strongly acidic dimetal site of this cation under mild conditions.

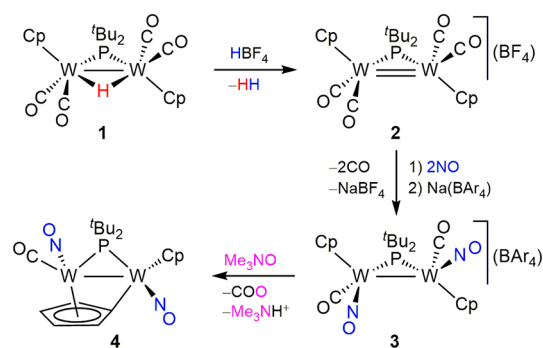
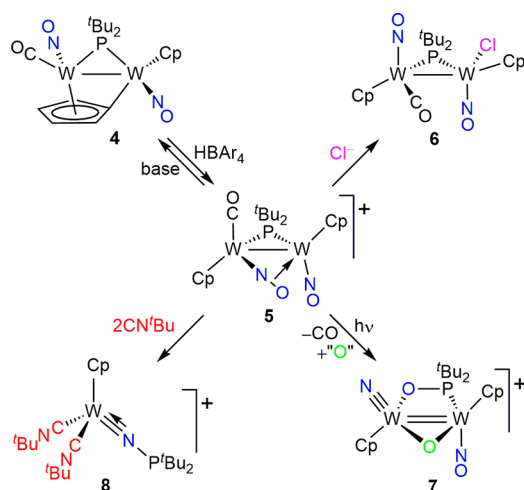
The mentioned $\kappa\eta$ nitrosyl-bridged complex was prepared through a multistep procedure starting from the corresponding hydrido-bridged complex $[W_2Cp_2(\mu-H)(\mu-P^tBu_2)(CO)_4]$ (1) (Schemes 1 and 2). The latter was prepared as reported previously for its PCy_2 -bridged analogue.¹⁰ This complex was

Received: June 24, 2022

Published: September 15, 2022



Scheme 1. Synthesis of Compound 4

Scheme 2. Reactivity of Compound 5^a

^aThe counterion is (BAr₄)[−] for all cations.

first dehydrogenated upon reaction with HBF₄·OEt₂, a process implemented previously for similar precursors,¹¹ to selectively give the corresponding unsaturated cationic complex [W₂Cp₂(μ-P^tBu₂)(CO)₄](BF₄) (2), which was not further purified. The latter was then reacted with NO (1 atm) at room temperature to give, after anion exchange with Na(BAr₄), the electron-precise dinitrosyl complex [W₂Cp₂(μ-P^tBu₂)(CO)₂(NO)₂](BAr₄) (3) in high yield. Spectroscopic data for this cation (see the Supporting Information, SI) were comparable to those recently reported by us for the PCy₂-bridged analogues,⁵ although now only the most symmetric isomer (with C₂ symmetry) is present in solution, likely because of the superior steric pressure imposed by the bulky P^tBu₂ ligand. Attempts to generate unsaturated derivatives of the latter cation via decarbonylation through irradiation with visible–UV light or prolonged heating in a refluxing toluene solution led only to the eventual decomposition of this complex. In contrast, decarbonylation took place easily even at 273 K upon reaction with Me₃NO (either hydrated or dehydrated) in a fluorobenzene solution, but then deprotonation of a cyclopentadienyl hydrogen by NMe₃ unexpectedly occurred too, to give the neutral cyclopentadienylidene-bridged derivative [W₂(μ-κ:η⁵-C₅H₄)Cp(μ-P^tBu₂)(CO)(NO)₂] (4) as a unique organometallic product, which was isolated in good yield (60%) after chromatographic workup. An X-ray study on the molybdenum analogue of 4 confirmed the coordination mode of the cyclopentadienylidene ligand in

these products,¹² as is also evident from the spectroscopic data of 4 (see the SI).

We have shown previously that bridging cyclopentadienylidene ligands can be protonated at the bridgehead carbon atom to regenerate cyclopentadienyl ligands,¹³ and this is also the case of compound 4. Indeed, the latter reacts instantaneously with [H(OEt₂)₂](BAr₄) in a dichloromethane solution to give the cyclopentadienyl derivative [W₂Cp₂(μ-P^tBu₂)(μ-κ:η-NO)(CO)(NO)](BAr₄) (5), in a process that can be reversed upon reaction with different bases (Scheme 2). The spectroscopic data for 5 denoted the presence of terminal CO and NO ligands, but the coordination mode of the second nitrosyl ligand was not obvious. An X-ray study revealed that the latter displays the unusual linear κ:η bridging mode (type D), with it being strongly bound to one metal via the nitrogen atom [W2–N3 = 1.823(4) Å], while π binding the second metal atom through its N–O bond [W1–N3 = 2.191(4) Å; W1–O3 = 2.138(3) Å], which then becomes significantly elongated [N3–O3 = 1.271(5) Å] and presumably debilitated, in agreement with its very low N–O stretching frequency (1366 cm^{−1}). The coordination mode of the NO ligand found in compound 5 was relatively unexpected because previous DFT studies on the decarbonylation products of the neutral complexes [W₂Cp₂(μ-PCy₂)(CO)₃(NO)] revealed that carbonyl is a better-suited ligand than NO for bridging two metal atoms in a linear κ:η fashion.¹⁴ In fact, density functional theory (DFT) calculations on the cation in 5 (Figure 1) and some possible isomers revealed that a linear

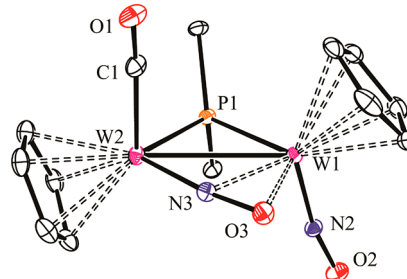


Figure 1. ORTEP diagram (30% probability) of the cation in 5, with hydrogen atoms and methyl groups omitted. Selected bond lengths (Å) and angles (deg): W1–W2 = 3.0978(3); W1–N2 = 1.771(4); W2–N3 = 1.823(4); W1–N3 = 2.191(4); W1–O3 = 2.138(3); N2–O2 = 1.205(5); N3–O3 = 1.271(5); W2–N3–O3 = 171.3(3).

κ:η-CO-bridged isomer would have the lowest energy, with the actual ordering found, when the nature and coordination mode of the bridging ligand are changed, being μ-κ:η-CO (0) < μ-κ:η-NO (+16) < μ-NO (+62) (relative Gibbs free energies at 298 K in kJ/mol; see the SI). Therefore, we conclude that 5 is a kinetic product. However, refluxing toluene solutions of this compound resulted in no detectable rearrangement.

Preliminary studies on the reactivity of 5 (Scheme 2) revealed that the π binding of the NO ligand can be removed upon the addition of simple donors, with concomitant rearrangement of the latter to terminal coordination, as found for κ:η-CO-bridged complexes. Thus, compound 5 reacts smoothly with CO (1 atm) at 348 K in a 1,2-dichloroethane solution to regenerate compound 3 selectively. It also easily adds chloride upon reaction with [N(PPh₃)₂]Cl at room temperature to give the neutral derivative [W₂ClCp₂(μ-P^tBu₂)(CO)(NO)₂] (6), with the chloride ligand in a cisoid positioning with respect to the phosphorus atom [P–W–Cl =

84.45(5)°, while the carbonyl ligand adopts a transoid arrangement [$\text{P}-\text{W}-\text{CO} = 105.4(2)^\circ$] with incipient semi-bridging geometry [$\text{W}-\text{W}-\text{CO} = 61.9(2)^\circ$; see the SI], also reflected in an anomalously low C–O stretching frequency in solution (1903 cm^{-1}). The addition of donor molecules, however, is limited by the high Brønsted acidity of the cation in **5**, which turned out to be deprotonated by conventional phosphines and amines or even upon dissolution of the complex in tetrahydrofuran, to give the cyclopentadienylidene precursor **4** quantitatively.

The carbonyl ligand in **5** can be removed photochemically at 243 K, in a process that also triggers cleavage of the N–O bond of one of the nitrosyl ligands and abstraction of an additional oxygen atom to give the oxo- and phosphinito-bridged nitrido complex $[\text{W}_2\text{Cp}_2(\text{N})(\mu\text{-O})(\mu\text{-OP}^t\text{Bu}_2)(\text{NO})](\text{BAR}_4)$ (**7**; Figure 2).¹⁵ Presumably, this N–O bond cleavage

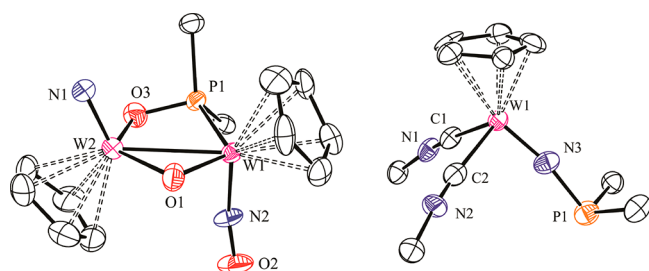


Figure 2. ORTEP diagrams (30% probability) of the cations in compounds **7** (left) and **8** (right), with hydrogen atoms and methyl groups omitted. Selected bond lengths (Å) and angles (deg) for **7**: $\text{W1}-\text{W2} = 2.823(1)$; $\text{W1}-\text{P1} = 2.47(1)$; $\text{W1}-\text{O1} = 1.97(2)$; $\text{W2}-\text{O1} = 2.29(3)$; $\text{W2}-\text{N1} = 1.70(2)$; $\text{W2}-\text{O3} = 1.98(2)$. Selected bond lengths (Å) and angles (deg) for **8**: $\text{W}-\text{C1} = 2.06(1)$; $\text{W}-\text{C2} = 2.12(2)$; $\text{W}-\text{N3} = 1.72(2)$; $\text{W}-\text{N3}-\text{P} = 170(1)$.

would follow decarbonylation to give first an oxo nitrido intermediate, as previously observed for some bridging nitrosyl ligands of type **A** at M_2^{7a} or $\text{M}_2\text{M}'$ centers ($\text{M} = \text{Mo}, \text{W}$; $\text{M}' = \text{Mn}, \text{Re}$).^{6,16} This would be followed by insertion of the oxo ligand into a W–P bond to generate the bridging phosphinito ligand and the addition of an extra oxygen atom (likely from trace dioxygen in the solution) to the unsaturated dimetal center thus generated.¹⁷

The exact nature and geometry of compound **7** could only be established through a crystallographic study, even if the precision in the geometrical parameters was modest, because the asymmetric cation lies on a plane of symmetry in the unit cell (a case of whole-body disorder; see the SI). The terminal nitrido ligand in the cation is strongly bound to a tungsten atom, as revealed by the short $\text{W2}-\text{N1}$ length of 1.70(2) Å¹⁸ and the presence of a relatively energetic W–N stretch of 962 cm^{-1} in the IR spectrum. The bridging oxo ligand displays W–O lengths above ca. 2.0 Å, which is indicative of modest π bonding.¹⁶ As a result, the cation can be considered to be a 32e complex, for which a double intermetallic bond should be proposed according to the 18e rule. This is consistent with the intermetallic distance of 2.823(1) Å in **7**, much shorter than those measured for the electron-precise complexes **5** and **6** [3.0978(3) and 3.1812(4) Å, respectively].

A different and unexpected N–O bond cleavage takes place in the reaction of **5** with CN^tBu . This reaction proceeds at 273 K, even when 1 equiv of isocyanide is used, to give, as the major phosphorus-containing species, the mononuclear bis-(isocyanide) complex $[\text{WCp}(\text{NP}^t\text{Bu}_2)(\text{CN}^t\text{Bu})_2](\text{BAR}_4)$ (**8**),

which displays a four-electron donor linear phosphinoimido ligand with a very short W–N separation of 1.72(2) Å (Figure 2).^{19–21} No intermediates were identified in this obviously multistep reaction, and further studies are now in progress to determine whether this reaction involves the oxidative addition of the $\kappa\text{:}\eta$ -bridging nitrosyl ligand to the dimetal center, as proposed for **7**, or rather follows from a direct deoxygenation reaction by isocyanide or carbonyl ligands, thus paralleling to some extent the reactions of the $\kappa\text{:}\kappa$ -nitrosyl-bridged complex $[\text{Mo}_2\text{Cp}_2(\mu\text{-PCy}_2)(\mu\text{-NO})(\text{NO})_2]$ with phosphites or CO.⁵

In summary, we have implemented an efficient preparative route for a new tungsten complex bearing a nitrosyl ligand in the rare linear $\kappa\text{:}\eta$ bridging coordination mode, by starting from the readily available P^tBu_2 -bridged precursor **1**, and involving the new complexes **2–4** as intermediate species. The reactivity of this nitrosyl-bridged complex indicates that π binding of the NO ligand to the second metal atom has two chemical effects not identified previously: (a) it facilitates the addition of ligands with concomitant rearrangement of the bridging nitrosyl into the terminal coordination mode, as found for linear $\kappa\text{:}\eta$ -CO-bridged complexes, and (b) it facilitates cleavage of the N–O bond of that ligand, possibly in two different ways: either through the oxidative addition of the ligand to the dimetal center or through deoxygenation by another ligand. None of these bond cleavage processes is known for the carbonyl ligand in the related $\kappa\text{:}\eta$ bridging mode.

■ ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/acs.inorgchem.2c02216>.

Preparative procedures, spectroscopic and microanalytical data for new compounds, crystallographic data, and results of DFT calculations (PDF)

Cartesian coordinates for all computed species (XYZ)

Accession Codes

CCDC 2181428–2181431 contain the supplementary crystallographic data for this paper. These data can be obtained free of charge via www.ccdc.cam.ac.uk/data_request/cif, or by emailing data_request@ccdc.cam.ac.uk, or by contacting The Cambridge Crystallographic Data Centre, 12 Union Road, Cambridge CB2 1EZ, UK; fax: +44 1223 336033.

■ AUTHOR INFORMATION

Corresponding Authors

Daniel García-Vivó – Departamento de Química Orgánica e Inorgánica/IUQOEM, Universidad de Oviedo, Oviedo E33071, Spain; orcid.org/0000-0002-2441-2486; Email: garciavdaniel@uniovi.es

Miguel A. Ruiz – Departamento de Química Orgánica e Inorgánica/IUQOEM, Universidad de Oviedo, Oviedo E33071, Spain; orcid.org/0000-0002-9016-4046; Email: marar@uniovi.es

Authors

M. Angeles Alvarez – Departamento de Química Orgánica e Inorgánica/IUQOEM, Universidad de Oviedo, Oviedo E33071, Spain; orcid.org/0000-0002-3313-1467

M. Esther García – Departamento de Química Orgánica e Inorgánica/IUQOEM, Universidad de Oviedo, Oviedo E33071, Spain; orcid.org/0000-0002-9185-0099

Ana M. Guerra – Departamento de Química Orgánica e Inorgánica/IUQOEM, Universidad de Oviedo, Oviedo E33071, Spain

Larry R. Falvello – Instituto de Nanociencia y Materiales de Aragón, Departamento de Química Inorgánica, CSIC, Universidad de Zaragoza, Zaragoza E-50009, Spain;

orcid.org/0000-0002-0444-996X

Complete contact information is available at:

<https://pubs.acs.org/10.1021/acs.inorgchem.2c02216>

Author Contributions

The manuscript was written through contributions of all authors.

Notes

The authors declare no competing financial interest.

ACKNOWLEDGMENTS

We thank Dr. Adrián Toyos for his inspiring work on PCy₂-bridged polynitrosyl complexes. We thank the MICINN of Spain and FEDER for financial support (Projects PGC2018-097366-B-I00 and PGC2018-093451-B-I00), the SCBI of the Universidad de Málaga, Spain, for access to computing facilities, and the X-ray units of Universidad de Oviedo and Universidad de Santiago de Compostela, Spain, for acquisition of the diffraction data.

REFERENCES

- (1) (a) Richter-Addo, G. B.; Legzdins, P. *Metal Nitrosyls*; Oxford University Press: Oxford, U.K., 1992. (b) Mingos, D. M. P., Ed. *Nitrosyl Complexes in Inorganic Chemistry, Biochemistry and Medicine: Structure and Bonding*; Springer: Berlin, 2014; Vols. 153 and 154.
- (2) For some recent reviews, see: (a) Lehnert, N.; Kim, E.; Dong, H. T.; Harland, J. B.; Hunt, A. P. H.; Manickas, E. C.; Oakley, K. M.; Pham, J.; Reed, G. C.; Alfaro, V. S. The Biologically Relevant Coordination Chemistry of Iron and Nitric Oxide: Electronic Structure and Reactivity. *Chem. Rev.* **2021**, *121*, 14682–14905. (b) Gallego, C. M.; Mazzeo, A.; Vargas, P.; Suárez, S.; Pellegrino, J.; Doctorovich, F. Azanone (HNO): generation, stabilization and detection. *Chem. Sci.* **2021**, *12*, 10410–10425. (c) Truzzi, D. R.; Medeiros, N. M.; Augusto, O.; Ford, P. C. Dinitrosyl Iron Complexes (DNICs): From Spontaneous Assembly to Biological Roles. *Inorg. Chem.* **2021**, *60*, 15835–15845. (d) Miranda, K. M.; Ridnour, L. A.; McGinity, C. L.; Bhattacharyya, D.; Wink, D. A. Nitric Oxide and Cancer: When to Give and When to Take Away. *Inorg. Chem.* **2021**, *60*, 15941–15947. (e) Ivanovic-Burmazovic, I.; Filipovic, M. R. Saying NO to H₂S: A story of HNO, HSNO and SSNO[−]. *Inorg. Chem.* **2019**, *58*, 4039–4051.
- (3) For some recent reviews, see: (a) Stepanenko, I.; Zalibera, M.; Schaniel, D.; Telsner, J.; Arion, V. B. Ruthenium-nitrosyl complexes as NO-releasing molecules, potential anticancer drugs, and photo-switches based on linkage isomerism. *Dalton Trans.* **2022**, *51*, 5367–5393. (b) Fraix, A.; Parisi, C.; Seggio, M.; Sortino, S. Nitric Oxide Photoreleasers with Fluorescent Reporting. *Chem. Eur. J.* **2021**, *27*, 12714–12725. (c) Muniz-Carvalho, E.; Silva-Sousa, E. H.; Bernardes-Génisson, V.; Gonzaga de Franca Lopes, L. When NO is not Enough: Chemical Systems, Advances and Challenges in the Development of NO and HNO Donors for Old and Current Medical Issues. *Eur. J. Inorg. Chem.* **2021**, *42*, 4316–4348. (d) Xiang, H. J.; Guo, M.; Liu, J. G. Transition-Metal Nitrosyls for Photocontrolled Nitric Oxide Delivery. *Eur. J. Inorg. Chem.* **2017**, *12*, 1586–1595.
- (4) For some recent reviews, see: (a) Han, L.; Cai, S.; Gao, M.; Hasegawa, J.; Wang, P.; Zhang, J.; Shi, L.; Zhang, D. Selective catalytic reduction of NO_x with NH₃ by using novel catalysts: State of the art and future prospects. *Chem. Rev.* **2019**, *119*, 10916–10976. (b) Lai, J.-K.; Wachs, I. E. A Perspective on the Selective Catalytic Reduction (SCR) of NO with NH₃ by Supported V₂O₅–WO₃/TiO₂ Catalysts. *ACS Catal.* **2018**, *8*, 6537–6551. (c) Liu, Y.; Zhao, J.; Lee, J.-M. Conventional and New Materials for Selective Catalytic Reduction (SCR) of NO_x. *ChemCatChem.* **2018**, *10*, 1499–1511.
- (5) (a) Alvarez, M. A.; García, M. E.; García-Vivó, D.; Ramos, A.; Ruiz, M. A.; Toyos, A. N–O Bond Activation and Cleavage Reactions of the Nitrosyl-Bridged Complexes [M₂Cp₂(μ-PCy₂)(μ-NO)(NO)₂] (M = Mo, W). *Inorg. Chem.* **2018**, *57*, 15314–15329. (b) Alvarez, M. A.; García, M. E.; García-Vivó, D.; Ruiz, M. A.; Toyos, A. Mild N–O Bond Cleavage Reactions of a Pyramidalized Nitrosyl Ligand Bridging a Dimolybdenum Center. *Inorg. Chem.* **2015**, *54*, 10536–10538.
- (6) Alvarez, M. A.; García, M. E.; Ruiz, M. A.; Toyos, A.; Vega, M. F. Low-Temperature N–O Bond Cleavage and Reversible N–P Bond Formation Processes in the Reactions of the Unsaturated Anions [M₂(η⁵-C₅H₅)₂(μ-PCy₂)(μ-CO)₂][−] (M = Mo, W) with the Nitrosyl Complex [Re(η⁵-C₅H₄Me)(CO)₂(NO)]⁺. *Inorg. Chem.* **2013**, *52*, 3942–3952.
- (7) For previous work from other groups on the N–O bond cleavage of nitrosyl ligands at binuclear complexes, see, for instance: (a) Legzdins, P.; Young, M. A.; Batchelor, R. J.; Einstein, F. W. B. Spontaneous Isomerization of Symmetric M(μ-NO)₂M Linkages to (ON)M:N:M:O Groupings. *J. Am. Chem. Soc.* **1995**, *117*, 8798–8806. (b) Böttcher, H.-C.; Graf, M.; Mereiter, K.; Kirchner, K. Reductive Dimerization of Nitric Oxide to *trans*-Hyponitrite in the Coordination Sphere of a Dinuclear Ruthenium Complex. *Organometallics* **2004**, *23*, 1269–1273. (c) Böttcher, H.-C.; Wagner, C.; Kirchner, K. Reaction Properties of the *trans*-Hyponitrite Complex [Ru₂(CO)₄(μ-H)(μ-PBu₂)₂](μ-Ph₂PCH₂PPh₂)(μ-η²-ONNO)]. *Inorg. Chem.* **2004**, *43*, 6294–6299. (d) Arikawa, Y.; Asayama, T.; Moriguchi, M.; Agari, S.; Onishi, M. Reversible N–N Coupling of NO Ligands on Dinuclear Ruthenium Complexes and Subsequent N₂O Evolution: Relevance to Nitric Oxide Reductase. *J. Am. Chem. Soc.* **2007**, *129*, 14160–14161. (e) Suzuki, T.; Tanaka, H.; Shiota, Y.; Sajith, P. K.; Arikawa, Y.; Yoshizawa, K. Proton-Assisted Mechanism of NO Reduction on a Dinuclear Ruthenium Complex. *Inorg. Chem.* **2015**, *54*, 7181–7191. (f) Arikawa, Y.; Hiura, J.; Tsuchii, C.; Kodama, M.; Matsumoto, N.; Umakoshi, K. Synthetic NO reduction cycle on a bis(pyrazolato)-bridged dinuclear ruthenium complex including photo-induced transformation. *Dalton Trans.* **2018**, *47*, 7399–7401. (g) Arikawa, Y.; Otsubo, Y.; Fujino, H.; Horiuchi, S.; Sakuda, E.; Umakoshi, K. Nitrite Reduction Cycle on a Dinuclear Ruthenium Complex Producing Ammonia. *J. Am. Chem. Soc.* **2018**, *140*, 842–847.
- (8) Type C: Ipaktschi, J.; Mohseni-Ala, J.; Dülmer, A.; Steffens, S.; Wittenburg, C.; Heck, J. Addition of Ynamines to the Tungsten η¹-Vinylidene Complexes (η⁵-C₅H₅)(NO)(CO)WCC(H)R. *Organometallics* **2004**, *23*, 4902–4909.
- (9) Type D: (a) Legzdins, P.; Rettig, S. J.; Veltheer, J. E. Organometallic nitrosyl chemistry. 52. A unique μ-η¹:η²-NO ligand in a bimetallic alkylidene nitrosyl complex of molybdenum. *J. Am. Chem. Soc.* **1992**, *114*, 6922–6923. (b) Legzdins, P.; Rettig, S. J.; Veltheer, J. E.; Batchelor, R. J.; Einstein, F. W. B. Thermal chemistry of molybdenum complex CpMo(NO)(CH₂CM₃)₂: Lewis base adducts of CpMo(NO)(:CHCM₃) and a new bonding mode for nitric oxide. *Organometallics* **1993**, *12*, 3575–3585.
- (10) García, M. E.; Riera, V.; Ruiz, M. A.; Rueda, M. T.; Sáez, D. Dimolybdenum and tungsten cyclopentadienyl carbonyls with electron-rich phosphido bridges. Synthesis of hydrido phosphido [M₂Cp₂(μ-H)(μ-PRR')(CO)₄] and unsaturated bisphosphido complexes [M₂Cp₂(μ-PR₂)(μ-PR'R'')(CO)_x] (x = 1, 2; R, R', R'' = Et, Cy, ^tBu). *Organometallics* **2002**, *21*, 5515–5525.
- (11) Alvarez, M. A.; García, M. E.; García-Vivó, D.; Ruiz, M. A.; Vega, M. F. Synthesis and reactivity of the triply bonded binuclear anion [W₂(η⁵-C₅H₅)₂(μ-PCy₂)(μ-CO)₂][−]: Tungsten makes a difference. *Organometallics* **2010**, *29*, 512–515.
- (12) For the X-ray-determined structure of [Mo₂(μ-κ⁵-C₅H₄)Cp(μ-P^tBu₂)(CO)(NO)₂], see: Alvarez, M. A.; García, M. E.; Guerra, A. M.; Ruiz, M. A. CSD Communication, 2022. DOI: 10.5517/ccdc.csd.cc2bskd3. CCDC 2168555. This complex could be prepared by following a synthetic route similar to that of **4**, but the yield of the last step was rather poor.

(13) Alvarez, M. A.; García, M. E.; Riera, V.; Ruiz, M. A. Protonation reactions on the binuclear complexes $[\text{W}_2\text{Cp}_2(\text{CO})_n(\mu\text{-L}_2)]$ ($n = 2, 4$). Chemical behavior of their hydrido and hydroxycarbonyl derivatives. *Organometallics* **1999**, *18*, 634–641.

(14) Alvarez, M. A.; García, M. E.; García-Vivó, D.; Rueda, M. T.; Ruiz, M. A.; Toyos, A.; Vega, M. F. Terminal vs. Bridging Coordination of CO and NO Ligands after Decarbonylation of $[\text{W}_2\text{Cp}_2(\mu\text{-PR}_2)(\text{CO})_3(\text{NO})]$ Complexes ($R = \text{Ph}, \text{Cy}$). An Experimental and Computational Study. *Dalton Trans.* **2017**, *46*, 10440–10451.

(15) Compound **7** was isolated through crystallization from the crude reaction mixture, but it was invariably contaminated with small amounts of the parent compound (typically ca. 15%; see the SI).

(16) García, M. E.; García-Vivó, D.; Melón, S.; Ruiz, M. A.; Graiff, C.; Tiripicchio, A. Low-Temperature N–O Bond Cleavage in Nitrosyl Ligands Induced by the Unsaturated Dimolybdenum Anion $[\text{Mo}_2(\eta^5\text{-C}_5\text{H}_5)_2(\mu\text{-PPh}_2)(\mu\text{-CO})_2]^-$. *Inorg. Chem.* **2009**, *48*, 9282–9293.

(17) Riera, V.; Ruiz, M. A.; Villafañe, F.; Bois, C.; Jeannin, Y. Reactivity of $(\mu\text{-CH}_2\text{PPh}_2)(\mu\text{-PPh}_2)\text{Mo}_2\text{Cp}_2(\text{CO})_2$ ($\text{Mo} = \text{Mo}$) toward iodine and chalcogens. Crystal structure of $(\mu\text{-O})[\text{CpMo}(\mu\text{-CH}_2\text{PPh}_2)(\mu\text{-O})(\mu\text{-OPPh}_2)\text{MoCp}(\text{CO})]_2$. *Organometallics* **1993**, *12*, 124–132.

(18) A search of the Cambridge Crystallographic Data Centre database (updated June 2022) on tungsten complexes bearing terminal nitrido ligands yielded just 15 complexes, almost all of them with W–N distances in the very narrow range of $1.68 \pm 0.01 \text{ \AA}$.

(19) Only a few complexes with terminal phosphinoimido ligands have been structurally characterized to date, none of them having Group 6 metal atoms. The observed W–N distance in **8** is significantly shorter than the corresponding distances measured for the niobium complexes $[\text{Nb}\{\text{N}(\text{Np})\text{Ar}\}_3(\text{NPA})]$ [$1.792(1) \text{ \AA}$] and $[\text{NbTp}^*(\text{Cl})(k^1\text{-CPhCEt}_2)\{\text{NP}(\text{N}^i\text{Pr}_2)_2\}]$ [$1.779(4) \text{ \AA}$],^{19,20} with the latter being the only organometallic phosphinoimido complex of this type structurally characterized so far.

(20) Eckhardt, A. K.; Riu, M. L. Y.; Müller, P.; Cummins, C. C. Staudinger Reactivity and Click Chemistry of Anthracene (A)-Based Azidophosphine N_3PA . *Inorg. Chem.* **2022**, *61*, 1270–1274.

(21) Etienne, M.; Mathieu, R.; Donnadieu, R. Reversible Migratory Insertion / β -Alkyl Elimination in α -Agostic Alkylniobium Alkyne Complexes. *J. Am. Chem. Soc.* **1997**, *119*, 3218–3228.

Recommended by ACS

Catalytic and Stoichiometric Reactions of the Parent Olefin Rhodium(I) Complex with Alkynes

Andrey V. Kolos, Dmitry S. Perekalin, *et al.*

OCTOBER 25, 2022
ORGANOMETALLICS

READ 

A Ruthenium Protic N-Heterocyclic Carbene Complex as a Precatalyst for the Efficient Transfer Hydrogenation of Aryl Ketones

Brian T. H. Tsui, Robert H. Morris, *et al.*

JULY 22, 2022
ORGANOMETALLICS

READ 

Carbene Transfer from a Pyridine Dipyrrolide Iron–Carbene Complex: Reversible Migration of a Diphenylcarbene Ligand into an Iron–Nitrogen Bond

Brett M. Hakey, Carsten Milsmann, *et al.*

AUGUST 01, 2022
ORGANOMETALLICS

READ 

Synthesis and Characterization of Rare-Earth Metal Complexes Bearing a 2-*N,N*-Dimethylamino-Ethylene-Imino-Functionalized Indolyl Ligand and Their Catalytic Activiti...

Yun Wei, Shaowu Wang, *et al.*

OCTOBER 18, 2022
ORGANOMETALLICS

READ 

Get More Suggestions >