

Self-propulsion of active colloids via ion release: theory and experiments

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We study the self-propulsion of a charged colloidal particle that releases ionic species using theory and experiments. We relax the assumptions of thin Debye-layers and weak-nonequilibrium effects assumed in classical phoretic models. This leads to a number of unexpected features that cannot be rationalized considering the classic phoretic framework: an active particle can reverse the direction of motion by increasing the rate of ions release and can propel even with zero surface charge. Our theory predicts that there are optimal conditions for self-propulsion and a novel regime in which the velocity is insensitive to the background electrolyte concentration. The theoretical results quantitatively captures the salt-dependent velocity measured in our experiments using active colloids that propel by decomposing urea via a surface enzymatic reaction.

To overcome random thermal fluctuations, microscopic active objects, such as bacteria [1] and molecular motors [2–4], convert chemical energy into motion. Inspired by nature, there has been an extensive effort in developing biomimetic synthetic active particles that employ catalytic chemical reactions to self-propel [5, 6]. Suspensions of these colloid-sized particles have been used as a model to investigate nonequilibrium active systems [7–10], ordering and phase transitions [11, 12] and collective communication [13–15]. Functionalized active colloids are also potential candidates for novel cancer therapies [16, 17], cargo transport [18] or other biomedical applications [19, 20]. In order to control the dynamics of active particles, be that for technological applications or to study their collective motion, one has to understand the mechanism that couples their chemical activity to motion.

Chemically active particles with a number of different propulsion mechanisms have been considered [21–27]. They rely on establishing a gradient of chemical potential along the surface of the particle, which drives its motion via phoretic flows [28]. Typically, the gradient of chemical potential is introduced by an asymmetric chemical activity that sustains gradients of solutes. For active col-

loids propelling by the generation of ionic species, the classic theory of phoresis [28, 29] has been widely applied [30–36]. However, the theory of phoresis was developed to study the motion of colloids in weak gradients of electrolytes [29] and thin Debye layers. It seems natural to ask whether the same framework is applicable to chemically active particles, which can be driven far from equilibrium or display a thick Debye layer.

To address this question, we investigate a general model for a charged active colloid that propels by releasing ionic species. We find that the surface flux of ions leads to a strong coupling between the solute transport and the charge balance in which the different diffusivity of positive and negative ions plays a fundamental role. Such coupling results in behaviors that cannot be accounted for using the standard phoretic models [28, 29], which break down for active colloids that release ions at a large rate. To demonstrate the relevance of the model, we compare its predictions with our measurements of the velocity of enzyme-powered colloids that propel through decomposition of urea into ammonia and carbon dioxide.

Consider a system containing a spherical non-conductive colloidal particle of radius R that is homogeneously charged with charge density q_s and is suspended

in an electrolyte solution. We assume that the electrolyte is a salt with unit valence and number density n_∞ , which is dissolved in a solvent with shear viscosity η and density ρ . The particle is chemically active and releases ions from a portion of its surface with a surface flux Q . The active portion of the particle surface is axisymmetric. We denote the excess number density of the cation with c_+ and the excess number density of the anion with c_- , with their diffusivity given by D_+ and D_- . See Figure 1 for an illustration. The model is applicable to self-propulsion mechanisms that involve surface ion flux such as chemical reactions involving electrolytes, dissolving salt particles or droplets absorbing ionic species. The asymmetric

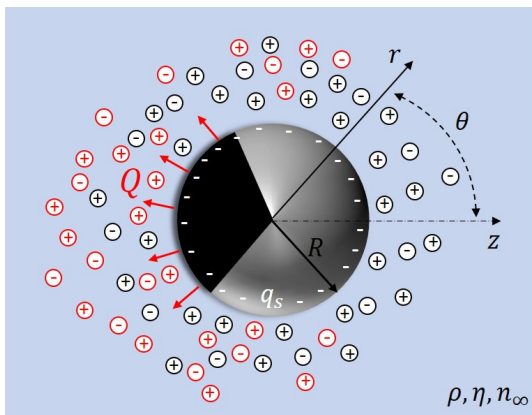


FIG. 1. A non-conductive particle with surface charge density q_s is suspended in an electrolyte solution, of number density n_∞ , and releases ions from a portion of its surface with a surface flux Q . The ions produced by the particle are displayed in red and the ions already in solution are displayed in black.

chemical activity of the particle drives its motion at a velocity $V\mathbf{e}_z$, with \mathbf{e}_z the unit vector along the z -axis. The ionic species released by the particle drive the system out of equilibrium introducing an additional charge density and electric field. The interaction between the charge density and the electric field results in an asymmetric electrostatic force density that propels the particle.

To evaluate the velocity of the active particle, we consider the balance of number density of the excess ionic species rescaled with n_∞ . Under the assumption that the species are dilute, that the system is at steady state and that advection is negligible, the dimensionless balance of number density reads:

$$\nabla^2 c_\pm \pm \nabla \cdot (c_\pm \nabla \Phi) = 0, \quad (1)$$

with the dimensionless electric potential Φ , which is scaled with the elementary charge e and the thermal energy $k_B T$. The excess ionic species vanish at infinity and are produced at the surface with a dimensionless flux $(\nabla c_+ + c_+ \nabla \Phi) \cdot \mathbf{n} = Q^* f(\theta)$ and $(\nabla c_- - c_- \nabla \Phi) \cdot \mathbf{n} = \beta^{-1} Q^* f(\theta)$, which applies at $r = 1$. $\beta = D_-/D_+$ denotes the ratio of diffusivity and $Q^* = QR/D_+ n_\infty$ controls the

ratio of the characteristic excess number density to the number density of the ionic species already in the solution. When the surface flux is small, $Q^* \ll 1$, the number density of the excess ionic species is much smaller than n_∞ and the system is close to equilibrium. Conversely, when $Q^* \gg 1$, the number density of the excess ionic species is much larger than n_∞ and the system is significantly driven out of equilibrium. The dimensionless function $f(\theta)$ determines the distribution of the surface chemical activity and can be decomposed in Legendre polynomials $f(\theta) = \sum_{l=0}^{\infty} f_l P_l(\cos \theta)$, with f_l its multipole moment.

The ionic species already dissolved in the solution are at equilibrium with respect to the electric potential. Their distribution is given by the Boltzmann distribution $n_\infty \exp(\pm e \Phi / k_B T)$. Therefore, the dimensionless balance of charge density in the fluid phase satisfies a modified Poisson-Boltzmann equation:

$$\nabla^2 \Phi = \frac{\kappa^2}{2} (c_- - c_+) + \kappa^2 \sinh(\Phi), \quad (2)$$

with κ the inverse of the Debye length, normalized by the radius of the particle, $\kappa = \sqrt{\frac{2e^2 n_\infty R^2}{k_B T \epsilon}}$, with the permittivity of the fluid ϵ . We consider zero electric potential at infinity, and a uniform surface charge density, $-\nabla \Phi \cdot \mathbf{n} = q_s^*$, with $q_s^* = q_s R e / k_B T \epsilon$.

For small Reynolds numbers, $Re = \rho V R / \eta \ll 1$, the velocity of the particle, made dimensionless with $e^2 \eta R / (k_B T)^2 \epsilon$, is calculated using the reciprocal theorem, which involves an integral over the fluid volume, Ω , outside the particle [37–39]:

$$V = -\frac{\kappa^2}{12\pi} \int_{\Omega} \mathbf{f} \cdot (\hat{\mathbf{v}} - \mathbf{e}_z) d\Omega, \quad (3)$$

with $\hat{\mathbf{v}}$ the velocity field around a sphere translating with unit velocity along \mathbf{e}_z and the electrostatic force density $\mathbf{f} = [c_+ - c_- - 2 \sinh(\Phi)] \nabla \Phi$. To evaluate Eq. (3) one needs to solve Eqs. (1)-(2). The propulsion velocity is a function of four dimensionless numbers: $V = V(q_s^*, Q^*, \kappa, \beta)$. We investigate next its functional dependence on the dimensionless numbers using asymptotic analysis and numerical simulations.

Particle velocity near equilibrium in the Debye-Hückel limit. — To investigate the mechanisms leading to self-propulsion at arbitrary κ , we evaluate the leading order velocity for $Q^* \ll 1$ and $q_s^* \ll 1$ using a regular perturbation expansion [39]. We consider the case of a monopolar and dipolar surface flux, $f_l = 0$ for $l > 1$. We find that the leading order velocity is

$$V = Q^{*2} \frac{f_0 f_1 (1 - \beta)^2}{\beta^2} g(\kappa) + Q^* q_s^* \frac{f_1 (1 - \beta)}{\beta} h(\kappa) + \text{h.o.t.}, \quad (4)$$

which is valid for arbitrary values of κ . The positive functions $g(\kappa)$ and $h(\kappa)$ encode the dependence on the dimen-

sionless Debye length and are reported in the supplementary material [39]. In the absence of the monopolar term $f_0 = 0$, Eq. (4) recovers the velocity found by Brown et al. [40]. In the case $\beta = 1$, the leading order terms given by Eq. (4) vanish, and the velocity is proportional to higher order terms. The well-known chemiophoretic contribution does not appear in Eq. (4) because it depends non-linearly on the surface charge density [29]. To assess the validity of the asymptotic theory, we tested Eq. (4) using numerical simulations [39].

The second term in the right hand side of Eq. (4) stems from the interaction of the charge density generated by the surface flux of ions with the equilibrium electric potential. It represents the standard phoretic contribution for small surface charge densities and predicts a velocity whose sign depends on $q_s^*(1 - \beta)$. For thin Debye layers, $\kappa \gg 1$, this term recovers the velocity predicted using the framework considered by Prieve et al. [29]. Conversely, the first term on the right hand side of Eq. (4) was overlooked by previous works [32, 35, 40] and shows that the particle propels even if not charged $q_s^* = 0$. This mechanism is fundamentally different from the standard phoretic mechanisms that rely on the existence of an interaction potential between the solute molecules and the surface of the particle [28, 29, 41]. Here, the motion of the particle originates entirely from the surface flux of ions, which couples transport of species and the balance of charges. This mechanism is only at play when $\beta \neq 1$, which highlights the importance of taking into account the different ionic diffusivities when the system is out of equilibrium. The first term in Eq. (4) is always positive and leads to motion towards the portion of the surface having the largest flux magnitude.

If the two terms in Eq. (4) have opposite signs, the direction of motion is determined by the prefactors and by the functions $g(\kappa)$ and $h(\kappa)$. For large Debye lengths, $\kappa \ll 1$, we find $g \propto \kappa^4$ and $h \propto \kappa^2$. In this limit, the second term in Eq. (4) dominates over the first one and the direction of motion is determined by $Q^*q_s^*(\beta - 1)$. In the case $\kappa \gg 1$, we find the asymptotic behavior $g(\kappa) \propto \kappa^{-1}$ and $h(\kappa) \propto \kappa^{-1}$, with the two functions having the same order of magnitude $\lim_{\kappa \rightarrow \infty} h(\kappa)/g(\kappa) = 2$ [39]. It is thus possible to reverse the direction of motion by changing the system parameters, e.g. salt concentration, particle size or surface flux, as observed in the experiments of Brown and Poon [42].

Numerical results for a Janus particle. — In many practical situations the surface flux of ionic species and the surface charge can be large and nonlinear effects become dominant. We investigate the velocity of a particle at non-vanishing Q^* and q_s^* using numerical simulations. We assume that half of the surface of the particle releases ionic species: $f(\theta) = 1$ for $\theta \leq \frac{\pi}{2}$, $f(\theta) = 0$ for $\theta > \frac{\pi}{2}$, and $Q^* > 0$.

Recent works [31, 34, 35] applied the framework of ionic phoresis developed by Prieve et al. [29], which as-

sumes a thin Debye-layer $\kappa \gg 1$, to study chemically active particles. It is not clear to what extent the theory developed for external gradients of ionic species can be applied to chemically active particles far from equilibrium. To address this point, we compare in Figure 2(a) the velocity predicted by Eqs. (1)-(3) to that predicted using the model by Prieve et al. [29, 39]. We consider the case $\beta = 2$ and the case of a thin Debye length $\kappa^2 = 1000$, where the framework of Prieve et al. [29] is expected to be valid. Figure 2(a) shows that the standard ionic phoretic model breaks down for systems driven sufficiently far from equilibrium $Q^* \approx 1$ and, for $Q^* \gg 1$, it can fail to predict even the correct direction of motion. We further discuss the deviation between our theory and that by Prieve et al. for $\kappa \gg 1$ but finite Q^* in the Supplementary note [39].

In Figures 2(b)-(c) we show that the functional dependence of V on Q^* can be divided in three regimes. We find a linear regime at small surface fluxes, $Q^* \ll 1$, where the standard phoretic mechanisms [28] are dominant and $V \propto Q^*$ with a slope that depends on q_s^* and on κ . In the case $Q^* \gg k^2$, we find a large flux regime for which the velocities at different κ collapse onto a master curve and become a function of $Q^*\kappa^2 = 2QR^3e^2/D_+k_B T\epsilon$. This dimensionless group is independent from the background electrolyte concentration n_∞ . In this regime, the number density of the excess ionic species is much larger than the background concentration in the solution and the propulsion velocity is independent of n_∞ . Further, in the limit $Q^*\kappa^2 \rightarrow \infty$, the velocity reaches a plateau that depends weakly on the surface charge. This is an interesting regime for practical applications because the propulsion velocity is insensitive to the background electrolyte concentration and to the surface charge of the particle. Most of the charge density in the fluid is generated by the excess ionic species and the surface charge density of the particle plays a minor role. At intermediate values of Q^* the magnitude of the velocity goes through a maximum, indicating that there is an optimal value of Q^* for self-propulsion, which depends on the system parameters. Finally, we find that the qualitative conclusions drawn from the asymptotic theory, Eq. (4), hold far beyond leading order. This is illustrated in Figure 2(b) where the velocity for $\beta = 2$ appears to change sign at a critical Q^* , as predicted by the Eq. (4) for the case $q_s^*(1 - \beta) < 0$. Conversely, in the case $q_s^*(1 - \beta) > 0$ displayed in Figure 2(c), the velocity is always positive as predicted by Eq. (4).

Comparison with experiments using enzyme-powered colloids. — We compare the prediction of the model with velocity measurements of chemically active colloidal particles, $R = 1 \mu\text{m}$. We perform experiments using functionalized particles that self-propel via the enzymatic decomposition of urea, $\text{CO}(\text{NH}_2)_2$, into ammonia, NH_3 , and carbon dioxide, CO_2 , catalyzed by urease tethered to their surface [39, 43]. The products form ions when dis-

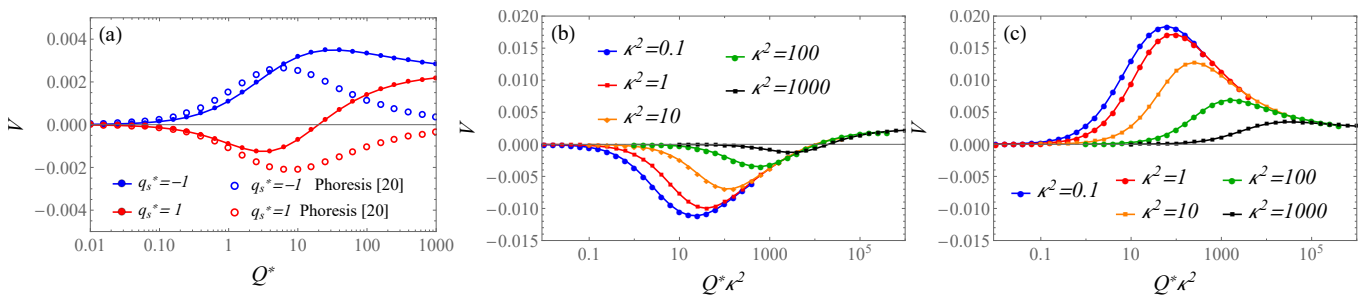


FIG. 2. (a) Velocity predicted by our theory compared to the predictions of the standard ionic phoretic model [29, 39] in the case $\kappa^2 = 1000$ and $\beta = 2$. (b) Velocity computed using simulations for the case $\beta = 2$ and $q_s^* = 1$. Since $(1 - \beta)q_s^*$ is negative, the velocity changes sign at a critical Q^* that depends on κ . (c) Dimensionless velocity computed using simulations for the case $\beta = 2$ and $q_s^* = -1$. Since $(1 - \beta)q_s^*$ is positive, the velocity does not change sign. For $Q^* \gg \kappa^2$ the velocity curves at different κ collapse onto a mastercurve. In this regime, the velocity is independent from the salt concentration.

solved in water according to the equilibria $\text{NH}_3 + \text{H}_2\text{O} \rightleftharpoons \text{NH}_4^+ + \text{OH}^-$ and $\text{CO}_2 + \text{H}_2\text{O} \rightleftharpoons \text{HCO}_3^- + \text{H}^+$. Recent works reported a local [44] and global [16] pH increase when urea is decomposed by the urease-powered colloids. This suggests that most of the ionic species are introduced by the reaction of NH_3 with water. To employ the theoretical framework developed here, we assume that the ammonia produced by urease turns instantaneously and irreversibly into NH_4^+ and OH^- and that these excess ionic species drive the motion of the particle. To test the theoretical predictions, we study the effect of the background salt concentration on the velocity of the particles.

The parameters of the model can be measured or estimated from the experimental conditions. The ratio of the diffusion coefficients of NH_4^+ and OH^- is known and yields $\beta = 2.7$ [45]. The velocity of the colloids is measured at 200 mM urea in deionized water and highly diluted phosphate-buffered saline solutions (PBS) with a maximum dilution ratio of 1 : 250. Since 90% of PBS is made of the monovalent electrolyte NaCl, we neglect the other electrolytes and assume that PBS contains NaCl only. We vary the background electrolyte concentration, n_∞ , over four orders of magnitude: from $n_\infty = 0.1 \mu\text{M}$, in deionized water, to $n_\infty = 5000 \mu\text{M}$ in 1 : 250 PBS solutions. Since the concentration of urea is much larger than the Michaelis-Menten constant, $K_m \approx 1 - 10 \text{mM}$ [46, 47], the enzymes work at their maximum rate and the reaction rate is independent from the local urea concentration. We estimate the surface flux of ionic species as $Q = k_{\text{cat}}\rho_{\text{ur}}$, with the turnover rate $k_{\text{cat}} \approx 10^4 \text{s}^{-1}$ [48, 49] and the surface density of enzymes $\rho_{\text{ur}} \approx 10^{16} \text{m}^{-2}$ [50], which yields $Q \approx 10^{20} \text{m}^{-2}\text{s}^{-1}$. The spherical symmetry of the surface flux is broken by the presence of imperfections of the particle surface [51] and by an asymmetric enzyme attachment to the silica surface that forms patches [50]. We assume that these features can be captured by a monopolar and dipolar surface flux $f_0 = f_1 = 1$ yielding a flux distribution $f(\theta) = 1 + \cos(\theta)$. Elec-

trophoretic measurements yield a surface charge density $q_s = -16 \mu\text{Cm}^{-2}$ [52]. The experiments range from a strongly nonequilibrium regime $Q^* \approx 3000$ and thick Debye layer $\kappa \approx 1$ in deionized water, to a near-equilibrium regime $Q^* \approx 0.5$ and a thin Debye layer $\kappa \approx 150$ in 1 : 250 PBS solutions [53].

The top panel of Figure 3 shows that there is a good agreement between the velocity measured in the experiments and that predicted by the model for a flux that is comparable to that expected from our estimate. Remarkably, at small n_∞ , the experiments confirm the existence of plateau where the velocity is independent from the background salt concentration of the medium as predicted by the theory, see Figure 2(b)-(c). Note that the plateau is observed in a regime where $Q^* \approx 3000$ and $\kappa \approx 1$ and therefore $Q^* \gg \kappa^2$ in which nonequilibrium effects are important. In the bottom panel of 3 we plot the axial component of the electrostatic force density, \mathbf{f} , around the active colloid. According to ionic phoretic models [28], the electrostatic force density should be confined to a thin layer next to the particle surface. In fact, Figures 3(a)-(b) show that far from equilibrium \mathbf{f} is spread around the particle. The break-down of phoretic models is particularly evident in the case of Figure 3(b) in which $\kappa \approx 20$ and one would expect \mathbf{f} to decay rapidly, which is not the case if $Q^* \gg 1$. As n_∞ is increased, the number density of the background ions becomes comparable to the excess ions released by the particle and the velocity decreases. In this regime, Figure 3(c) shows that the electrostatic force density is concentrated within a thin boundary layer as assumed in phoretic models.

Conclusions. — We have shown that the nonequilibrium nature of active colloids that release ions can lead to qualitatively different behaviors compared to standard phoretic mechanisms. The features we identified cannot be explained using the classic ionic phoretic approaches, which break down if the particle is significantly driven out of equilibrium by its chemical activity, $Q^* \approx 1$. The predictions quantitatively explain the dependence of the ve-

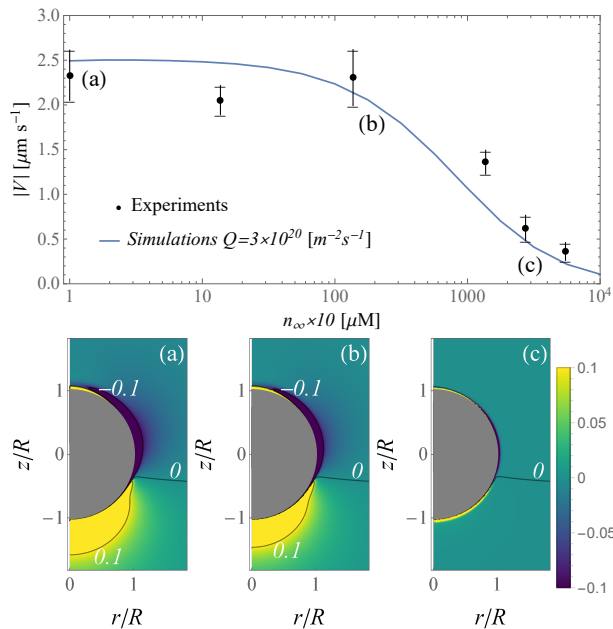


FIG. 3. Effect of the background electrolyte concentration on the particle velocity. Top panel: comparison between the magnitude of the velocity predicted by our theory and our measurements performed on catalytic active particles propelling via enzymatic decomposition of urea. Panels (a), (b) and (c): the z-component of the dimensionless electrostatic force density, f , corresponding to the cases marked in the top panel.

locity of urease-powered colloids on the background concentration of electrolytes measured in our experiments, thus confirming that the regimes investigated here are relevant for experiments in electrolyte solutions. The effects discovered here shed light on the mechanisms behind the propulsion of active colloids far from equilibrium. Exploiting the new phenomena discussed could open new possibilities for engineering their cooperative motion and for tuning their interactions with confining surfaces.

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- [51] *See Figure 3 of supplementary material.*
- [52] *See Figure 4 of supplementary material.*
- [53] *See Figure 5 of supplementary material.*